Name: _____ Class: _____ Date: _____ ID: A

AP Chemistry Chapter 4 Practice Exam

D 1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

$2H_2S + 3O_2 \rightarrow 2SO_2 + 2H_2O$

$SO_2 + 2H_2S \rightarrow 3S + 2H_2O$

How much sulfur (in grams) is produced from 21.5 grams of O_2?

A) 97.0 g
B) 14.4 g
C) 21.5 g
D) 43.1 g
E) none of these

$\frac{21.5 g\ O_2}{} \left| \frac{1 mol\ O_2}{32.00 g\ O_2} \right| \frac{2 SO_2}{3 O_2} \left| \frac{3 S}{SO_2} \right| \frac{32.06 g\ S}{1 mol\ S} = 43.1 g\ S$

B 2. Iron is produced from its ore by the reactions:

$2C(s) + O_2(g) \rightarrow 2CO(g)$

$Fe_2O_3(s) + 3CO(g) \rightarrow 2Fe(s) + 3CO_2(g)$

How many moles of $O_2(g)$ are needed to produce 4.6 moles of $Fe(s)$?

A) 2.3 mole O_2
B) 3.5 mole O_2
C) 4.6 mole O_2
D) 6.9 mole O_2
E) 13.8 mole O_2

$\frac{4.6 mol\ Fe(s)}{} \left| \frac{3 CO}{2 Fe} \right| \frac{1 O_2}{2 CO} = 3.45 mol\ O_2$

B 3. Suppose the reaction $Ca_3(PO_4)_2 + 3H_2SO_4 \rightarrow 3CaSO_4 + 2H_3PO_4$ is carried out starting with 153 g of $Ca_3(PO_4)_2$ and 76.8 g of H_2SO_4. How much phosphoric acid will be produced?

A) 76.7 g
B) 51.1 g
C) 229.8 g
D) 115.1 g
E) 96.7 g

$\frac{153 g\ Ca_3(PO_4)_2}{} \left| \frac{1 mol\ Ca_3(PO_4)_2}{} \right| \frac{2 H_3PO_4}{1 Ca_3(PO_4)_2} \left| \frac{}{1 mol\ H_3PO_4} \right. =$

$\frac{76.8 g\ H_2SO_4}{} \left| \frac{1 mol\ H_2SO_4}{64 g\ H_2SO_4} \right| \frac{2 H_3PO_4}{3 H_2SO_4} \left| \frac{}{1 mol\ H_3PO_4} \right. =$

B 4. SO_2 reacts with H_2S as follows:

$2H_2S + SO_2 \rightarrow 3S + 2H_2O$

When 7.50 g of H_2S reacts with 12.75 g of SO_2, which statement applies?

A) 6.38 g of sulfur are formed.
B) 10.6 g of sulfur are formed.
C) 0.0216 moles of H_2S remain.
D) 1.13 g of H_2S remain.
E) SO_2 is the limiting reagent.

$\frac{7.50 g\ H_2S}{} \left| \frac{1 mol\ H_2S}{34.03 g\ H_2S} \right| \frac{3 S}{2 H_2S} \left| \frac{32 g\ S}{1 mol\ S} \right. = 10.87$

$\frac{12.75 g\ SO_2}{} \left| \frac{1 mol\ SO_2}{64 g\ SO_2} \right| \frac{3 S}{1 SO_2} \left| \frac{32 g\ S}{1 mol\ S} \right. = 19.13$

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F Rizvi

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