CHAPTER 9 — STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	pi bond

When sharing of electrons occurs, the attraction between atoms that results is called a
(1) __covalent bond__ . When two or more bond through dissociation
energy is released, and the process in (2) __exothermic__ . When two or more
atoms bond by means of electron sharing the resulting particle is called a(n)
(3) __molecule__ . If the electrons shared are centered between the two atoms, the
attraction is called a(n) (4) __sigma bond__ . If the sharing involves the overlap of
parallel orbitals, the attachment is called a(n) (5) __pi bond__ .

In your textbook, read about simple and multiple bonds and bond strength.

Circle the letter of the choice that best completes the statement or answers the question.

6. In what form do elements exist in hydrogen, nitrogen, and oxygen normally occur?
a. as single atoms
b. as molecules containing two atoms
c. as molecules containing three atoms
d. as molecules containing four atoms

7. How many electrons are shared in a double covalent bond?
a. one
b. two
c. two
d. four

8. Bond length is the distance between
a. two molecules of the same substance
b. the electrons in two attached atoms
c. the nuclei of two attached atoms
d. the orbitals of two attached atoms

9. Which of the following relationships relating to bond length is generally correct?
a. the shorter the bond, the stronger the bond
b. the shorter the bond, the weaker the bond
c. the shorter the bond, the less the electrons in it
d. the shorter the bond, the lower the bond dissociation energy

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CHAPTER 9 — STUDY GUIDE FOR CONTENT MASTERY

Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

___false___ **1.** Binary molecular compounds are generally composed of a metal and a nonmetal.

___false___ **2.** The second element in the formula of a binary compound is named using the suffix -ide.

___false___ **3.** The prefix tetra- indicates three atoms.

___true___ **4.** The prefix hexa- indicates six atoms.

___true___ **5.** In naming the first element in a formula, the prefix mono- is not used.

___true___ **6.** For binary acids, the hydrogen part of the compound is named using the prefix hydro-.

___false___ **7.** An oxyacid contains only two elements.

___false___ **8.** If the name of the anion of an oxyacid ends in -ite, the acid name contains the suffix -ous.

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A		Column B
c	**9.** CO	a. hydrochloric acid
j	**10.** CO_2	b. dinitrogen tetroxide
g	**11.** H_2CO_3	c. carbon monoxide
e	**12.** NH_3	d. nitrous acid
b	**13.** N_2O_4	e. ammonia
d	**14.** HNO_2	f. nitric acid
f	**15.** HNO_3	g. carbonic acid
a	**16.** HCl	h. carbon dioxide
h	**17.** $HClO_3$	i. carbon dioxide

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Neil Thomson

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