

The Periodic Table and Periodic Law
Study Guide for Content Mastery

Name:
Date:
Period:

Section 6.1: Development of the Modern Periodic Table

In your textbook, reads about the *history of the periodic table's development*.

Use each of the terms below just once to complete the passage.

| | | | | |
|------------------|---------------|---------------|---------|--------------|
| Octaves | Atomic mass | atomic number | nine | elements |
| Properties | Henry Moseley | eight | protons | periodic law |
| Dmitri Mendeliev | accepted | | | |

The table below was developed by John Newlands and is based on a relationship called the law of (1) _____. According to this law, the properties of the elements repeated every (2) _____ elements. Thus, for example, element two and element (3) _____ have similar properties. The law of octaves did not work for all known elements and was not generally (4) _____.

| | | | | | | |
|---|----|----|----|----|----|----|
| 1 | 2 | 3 | 4 | 5 | 6 | 7 |
| H | Li | G | Bo | C | N | O |
| 8 | 9 | 10 | 11 | 12 | 13 | 14 |
| F | Na | Mg | Al | Si | P | S |

The first periodic table is mostly credited to (5) _____. In his table, the elements were arranged according to increasing (6) _____. One important result of this table was that the existence and properties of undiscovered (7) _____ could be predicted.

The elements in the modern periodic table are arranged according to increasing (8) _____, as a result of the work of (9) _____. This arrangement is based on number of (10) _____ in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) _____, which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) _____.

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A Loxley



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