



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or covalent.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - H_2O_2
- Write the formula for the following compounds:
 - hydrogen chloride
 - hydrogen sulfide
 - hydrogen cyanide
 - hydrogen peroxide
 - hydrogen selenide
 - hydrogen telluride

- Write the molecular formula for each of the following compounds:
 - hydrogen cyanide
 - hydrogen selenide
 - hydrogen telluride
 - hydrogen peroxide
 - hydrogen sulfide
 - hydrogen chloride
- Write the molecular formula for each of the following compounds:
 - hydrogen cyanide
 - hydrogen selenide
 - hydrogen telluride
 - hydrogen peroxide
 - hydrogen sulfide
 - hydrogen chloride

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride: $\text{H}-\text{F}$
- Draw the Lewis structure for hydrogen peroxide: $\text{H}-\text{O}-\text{O}-\text{H}$
- Draw the Lewis structure for carbon dioxide: $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for ethane: $\text{H}-\text{C}-\text{C}-\text{H}$
- Draw the Lewis structure for acetylene: $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for hydrogen cyanide: $\text{H}-\text{C}\equiv\text{N}$



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by σ_{1s} , σ_{2s} , σ_{2p} , σ_{3s} , σ_{3p} , σ_{4s} , σ_{4p} , σ_{5s} , σ_{5p} , σ_{6s} , σ_{6p} , σ_{7s} , σ_{7p} , σ_{8s} , σ_{8p} , σ_{9s} , σ_{9p} , σ_{10s} , σ_{10p} , σ_{11s} , σ_{11p} , σ_{12s} , σ_{12p} , σ_{13s} , σ_{13p} , σ_{14s} , σ_{14p} , σ_{15s} , σ_{15p} , σ_{16s} , σ_{16p} , σ_{17s} , σ_{17p} , σ_{18s} , σ_{18p} , σ_{19s} , σ_{19p} , σ_{20s} , σ_{20p} , σ_{21s} , σ_{21p} , σ_{22s} , σ_{22p} , σ_{23s} , σ_{23p} , σ_{24s} , σ_{24p} , σ_{25s} , σ_{25p} , σ_{26s} , σ_{26p} , σ_{27s} , σ_{27p} , σ_{28s} , σ_{28p} , σ_{29s} , σ_{29p} , σ_{30s} , σ_{30p} , σ_{31s} , σ_{31p} , σ_{32s} , σ_{32p} , σ_{33s} , σ_{33p} , σ_{34s} , σ_{34p} , σ_{35s} , σ_{35p} , σ_{36s} , σ_{36p} , σ_{37s} , σ_{37p} , σ_{38s} , σ_{38p} , σ_{39s} , σ_{39p} , σ_{40s} , σ_{40p} , σ_{41s} , σ_{41p} , σ_{42s} , σ_{42p} , σ_{43s} , σ_{43p} , σ_{44s} , σ_{44p} , σ_{45s} , σ_{45p} , σ_{46s} , σ_{46p} , σ_{47s} , σ_{47p} , σ_{48s} , σ_{48p} , σ_{49s} , σ_{49p} , σ_{50s} , σ_{50p} , σ_{51s} , σ_{51p} , σ_{52s} , σ_{52p} , σ_{53s} , σ_{53p} , σ_{54s} , σ_{54p} , σ_{55s} , σ_{55p} , σ_{56s} , σ_{56p} , σ_{57s} , σ_{57p} , σ_{58s} , σ_{58p} , σ_{59s} , σ_{59p} , σ_{60s} , σ_{60p} , σ_{61s} , σ_{61p} , σ_{62s} , σ_{62p} , σ_{63s} , σ_{63p} , σ_{64s} , σ_{64p} , σ_{65s} , σ_{65p} , σ_{66s} , σ_{66p} , σ_{67s} , σ_{67p} , σ_{68s} , σ_{68p} , σ_{69s} , σ_{69p} , σ_{70s} , σ_{70p} , σ_{71s} , σ_{71p} , σ_{72s} , σ_{72p} , σ_{73s} , σ_{73p} , σ_{74s} , σ_{74p} , σ_{75s} , σ_{75p} , σ_{76s} , σ_{76p} , σ_{77s} , σ_{77p} , σ_{78s} , σ_{78p} , σ_{79s} , σ_{79p} , σ_{80s} , σ_{80p} , σ_{81s} , σ_{81p} , σ_{82s} , σ_{82p} , σ_{83s} , σ_{83p} , σ_{84s} , σ_{84p} , σ_{85s} , σ_{85p} , σ_{86s} , σ_{86p} , σ_{87s} , σ_{87p} , σ_{88s} , σ_{88p} , σ_{89s} , σ_{89p} , σ_{90s} , σ_{90p} , σ_{91s} , σ_{91p} , σ_{92s} , σ_{92p} , σ_{93s} , σ_{93p} , σ_{94s} , σ_{94p} , σ_{95s} , σ_{95p} , σ_{96s} , σ_{96p} , σ_{97s} , σ_{97p} , σ_{98s} , σ_{98p} , σ_{99s} , σ_{99p} , σ_{100s} , σ_{100p}

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Camilla Rothe



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