

Chapter 8 Homework

1. If $\Delta H = -54.74 \text{ kJ}$ for a chemical reaction, that reaction _____.
- A. is exothermic
 - B. is endothermic
 - C. cannot occur
 - D. none of the above
2. A _____ is an experimental set-up that can be used to measure the heat of a chemical reaction or a physical change and determine the specific heat capacity of a substance.
- A. manometer
 - B. thermometer
 - C. calorimeter
 - D. whole-body calorimetry
3. You are asked to design an experiment to measure the enthalpy of a given reaction. The reaction needs to be at atmospheric pressure so which of the following techniques would be ideal for this experiment?
- A. bomb manometry
 - B. bomb calorimetry
 - C. coffee-cup calorimetry
 - D. whole-body calorimetry
4. Given the ΔH°_f , calculate the $\Delta H^\circ_{\text{rxn}}$ for the formation of diamond from graphite.

$$\Delta H^\circ_f (\text{C, graphite}) = 0 \text{ kJ/mol}$$

$$\Delta H^\circ_f (\text{C, diamond}) = 2 \text{ kJ/mol}$$

A. 0 kJ/mol

B. +2 kJ/mol

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