

## Chapter 9 Practice Problems: Chemical Names and Formulas

### Section 9.1: Naming Ions

- What is the charge on the ion typically formed by each element?
  - oxygen
  - iodine
  - sodium
  - aluminum
  - nickel, two electrons lost
  - magnesium
- How many electrons does the neutral atom gain or lose when each ion forms?
  - $\text{Cr}^{3+}$
  - $\text{P}^{3-}$
  - $\text{Li}^+$
  - $\text{Ca}^{2+}$
  - $\text{Cl}^-$
  - $\text{O}^{2-}$
- Name each ion. Identify each as a cation or an anion.
  - $\text{Sn}^{2+}$
  - $\text{Co}^{3+}$
  - $\text{Br}^-$
  - $\text{K}^+$
  - $\text{H}^-$
  - $\text{Mn}^{2+}$
- Write the formula (including the charge) for each ion. Use table 9.3 as if necessary.
  - carbonate ion
  - nitrite ion
  - sulfate ion
  - hydroxide ion
  - chromate ion
  - ammonium ion
- Name each of the following ions. Identify each as a cation or an anion.
  - $\text{CN}^-$
  - $\text{HCO}_3^-$
  - $\text{PO}_4^{3-}$
  - $\text{Cl}^-$
  - $\text{Ca}^{2+}$
  - $\text{SO}_3^{2-}$

### 9.2 Naming and Writing Formulas for Ionic Compounds:

- Write the formulas for these binary ionic compounds.
  - magnesium oxide
  - tin(II) fluoride
  - potassium iodide
  - aluminum chloride
  - sodium sulfide
  - ferric bromide
- Write the formulas for the compounds formed from these pairs of ions.
  - $\text{Ba}^{2+}$  and  $\text{Cl}^-$
  - $\text{Ag}^+$  and  $\text{I}^-$
  - $\text{Ca}^{2+}$  and  $\text{S}^{2-}$
  - $\text{K}^+$  and  $\text{Br}^-$
  - $\text{Al}^{3+}$  and  $\text{O}^{2-}$
  - $\text{Fe}^{2+}$  and  $\text{O}^{2-}$
- Name the following binary ionic compounds.
  - $\text{MnO}_2$
  - $\text{Li}_3\text{N}$
  - $\text{CaCl}_2$
  - $\text{SrBr}_2$
  - $\text{NiCl}_2$
  - $\text{K}_2\text{S}$
  - $\text{Fe}_3\text{N}_2$
  - $\text{SnCl}_4$
- Write formulas for the following ionic compounds.
  - sodium phosphate
  - magnesium sulfate
  - sodium hydroxide
  - potassium cyanide
  - ammonium chloride
  - potassium dichromate
- Write formulas for the following ionic compounds.
  - $\text{NH}_4^+$  and  $\text{SO}_4^{2-}$
  - $\text{K}^+$  and  $\text{NO}_3^-$
  - barium ion and hydroxide ion
  - lithium ion and carbonate ion

# Chapter 9 Chemical Names And Formulas Practice Problems Answers

**Arturo Cuomo**



**Chapter 9 Chemical Names And Formulas Practice Problems Answers:**

## Unveiling the Power of Verbal Art: An Emotional Sojourn through **Chapter 9 Chemical Names And Formulas Practice Problems Answers**

In a world inundated with screens and the cacophony of immediate interaction, the profound energy and psychological resonance of verbal artistry frequently disappear into obscurity, eclipsed by the continuous onslaught of noise and distractions. Yet, set within the lyrical pages of **Chapter 9 Chemical Names And Formulas Practice Problems Answers**, a captivating perform of fictional elegance that impulses with raw feelings, lies an remarkable trip waiting to be embarked upon. Composed with a virtuoso wordsmith, this interesting opus books readers on an emotional odyssey, gently exposing the latent potential and profound influence stuck within the elaborate internet of language. Within the heart-wrenching expanse of this evocative evaluation, we shall embark upon an introspective exploration of the book is central themes, dissect their captivating publishing style, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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