

## Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
  - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
  - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
  - $\frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} = \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} = \frac{1 \text{ mol CaO}}{1 \text{ mol Ca}}$  Remember the ratios can be reduced as well!!!
  - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$   $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2}$   $\frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol HF}}{1 \text{ mol MgF}_2}$   $\frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol HF}}{1 \text{ mol MgF}_2}$   $\frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol HF}}{1 \text{ mol MgF}_2}$
- $\frac{105.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{105.99 \text{ g Na}_2\text{CO}_3} = \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} = \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2}$   $\frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} \cdot \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} = \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3}$
- Molar Mass
  - Definition: the mass of one mole of a substance
  - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen  $\rightarrow$  Water
  - 5.0 mol H<sub>2</sub>
  - 2.5 mol O<sub>2</sub>
- Ethane combustion reaction
  - 15.8 mol O<sub>2</sub>
  - 9.00 mol CO<sub>2</sub> & 13.5 mol H<sub>2</sub>O
- 575 g Na & 886 g Cl<sub>2</sub>
- Iron production reaction
  - 75.1 mol CO
  - 50.1 mol Fe & 75.1 mol CO<sub>2</sub>
- 87,390 g CO or  $8.739 \times 10^4 \text{ g CO}$  & 12,610 g H<sub>2</sub> or  $1.261 \times 10^4 \text{ g H}_2$
- Lightning reaction of N<sub>2</sub> & O<sub>2</sub>
  - 1,100 g NO<sub>2</sub> or  $1.10 \times 10^3 \text{ g NO}_2$
  720. g NO or  $7.20 \times 10^2 \text{ g NO}$
- CO<sub>2</sub> removal in spacecraft
  - 42.04 mol NaOH
  - 21.02 mol Na<sub>2</sub>CO<sub>3</sub> & 21.02 mol H<sub>2</sub>O
- Double replacement AgNO<sub>3</sub> & NaBr
  - 463 g NaBr
  - 845 g AgBr
- Soda-acid fire extinguisher
  - 3.058 mol NaHCO<sub>3</sub>
  - 3.058 mol CO<sub>2</sub>, 1.529 mol Na<sub>2</sub>SO<sub>4</sub>, 3.058 mol H<sub>2</sub>O
- Sulfuric acid-sodium hydroxide neutralization reaction
  - H<sub>2</sub>SO<sub>4</sub> + 2 NaOH  $\rightarrow$  Na<sub>2</sub>SO<sub>4</sub> + 2 H<sub>2</sub>O
  - 37 g H<sub>2</sub>SO<sub>4</sub>
  - 53 g Na<sub>2</sub>SO<sub>4</sub> & 14 g H<sub>2</sub>O
- Copper & silver nitrate reaction
  - 0.0104 mol Cu(NO<sub>3</sub>)<sub>2</sub>
  - 0.0104 mol Cu & 0.0209 mol AgNO<sub>3</sub>
- Aspirin production
  - 13.5 kg C<sub>9</sub>H<sub>8</sub>O<sub>4</sub>
  - 7.66 kg C<sub>9</sub>H<sub>8</sub>O<sub>3</sub>
  - 4.29 L HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

# Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers

**Scott C. Dulebohn**



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web solution from the laws of motion in the case of rotational motion we know that  $\omega = \omega_0 + \alpha t$  from the given data  $\omega_0 = 900 \text{ rpm}$   $900 \times 2\pi / 60 \text{ rad s}$   $\omega = 0$  and  $t = 60 \text{ s}$  hence angular deceleration  $\alpha = -\pi / 2$  question 3 a bus is moving in a circular track of radius 1000cm with a speed of 10m s

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web circular motion examples the motion of a car on a level road let us assume a car is moving on a level road with mass  $m$  and  $g$  as the acceleration due to gravity acting upon it when the car is taking a turn on the road a total of three forces act simultaneously on the car the weight of the car  $mg$

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web inertia and circular motion how can you explain the sensation of there being an outward force when you move along a curved path in a car or an amusement park ride a commonly held belief by beginning physics students is that objects moving in circles experience a centrifugal outward force

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web newton s first law of motion states the following a body at rest tends to remain at rest a body in motion tends to remain in motion at a constant velocity unless acted on by a net external force recall that constant velocity means that the body moves in a straight line and at a constant speed

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web the simplest case of circular motion is uniform circular motion where an object travels a circular path at a constant speed note that unlike speed the linear velocity of an object in circular motion is constantly changing because it is always changing direction

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