

## 10

## CHEMICAL QUANTITIES

## Chapter Test A

## A. Matching

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

## Column A

- \_\_\_\_\_ 1. standard temperature and pressure
- \_\_\_\_\_ 2. molar volume
- \_\_\_\_\_ 3. representative particle
- \_\_\_\_\_ 4. molar mass
- \_\_\_\_\_ 5. atomic mass
- \_\_\_\_\_ 6. Avogadro's number
- \_\_\_\_\_ 7. mole (mol)
- \_\_\_\_\_ 8. empirical formula
- \_\_\_\_\_ 9. percent composition
- \_\_\_\_\_ 10. molar mass

## Column B

- a. the volume occupied by a mole of any gas at STP (22.4 L)
- b. the mass in amu of an atom of an element
- c. the number of representative particles in a mole of a substance
- d. the percent by mass of each element in a compound
- e. the mass of a mole of a substance
- f. the formula that gives the lowest whole-number ratio of the elements in a compound
- g. an atom, a formula unit, or a molecule
- h. the mass in grams of a mole of any pure substance
- i.  $6.02 \times 10^{23}$  representative particles of a substance
- j.  $0^\circ\text{C}$  and 101.3 kPa

## B. Multiple Choice

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol,  $\text{C}_3\text{H}_8\text{O}$ ?
- a.  $4 \times (6.02 \times 10^{23})$
- b. 4
- c. 32
- d.  $32 \times (6.02 \times 10^{23})$
- \_\_\_\_\_ 12. The mass of a mole of NaCl is its:
- a. molar mass.
- b. empirical formula.
- c. gram atomic mass.
- d. atomic mass.

# Chemical Quantities Chapter 10 Test Answer Key

**VM Jensen**



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## **Table of Contents Chemical Quantities Chapter 10 Test Answer Key**

1. Understanding the eBook Chemical Quantities Chapter 10 Test Answer Key
  - The Rise of Digital Reading Chemical Quantities Chapter 10 Test Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Quantities Chapter 10 Test Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Quantities Chapter 10 Test Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Quantities Chapter 10 Test Answer Key
  - Personalized Recommendations
  - Chemical Quantities Chapter 10 Test Answer Key User Reviews and Ratings
  - Chemical Quantities Chapter 10 Test Answer Key and Bestseller Lists
5. Accessing Chemical Quantities Chapter 10 Test Answer Key Free and Paid eBooks
  - Chemical Quantities Chapter 10 Test Answer Key Public Domain eBooks
  - Chemical Quantities Chapter 10 Test Answer Key eBook Subscription Services
  - Chemical Quantities Chapter 10 Test Answer Key Budget-Friendly Options

6. Navigating Chemical Quantities Chapter 10 Test Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Quantities Chapter 10 Test Answer Key Compatibility with Devices
  - Chemical Quantities Chapter 10 Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Quantities Chapter 10 Test Answer Key
  - Highlighting and Note-Taking Chemical Quantities Chapter 10 Test Answer Key
  - Interactive Elements Chemical Quantities Chapter 10 Test Answer Key
8. Staying Engaged with Chemical Quantities Chapter 10 Test Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Quantities Chapter 10 Test Answer Key
9. Balancing eBooks and Physical Books Chemical Quantities Chapter 10 Test Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Quantities Chapter 10 Test Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemical Quantities Chapter 10 Test Answer Key
  - Setting Reading Goals Chemical Quantities Chapter 10 Test Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Quantities Chapter 10 Test Answer Key
  - Fact-Checking eBook Content of Chemical Quantities Chapter 10 Test Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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