

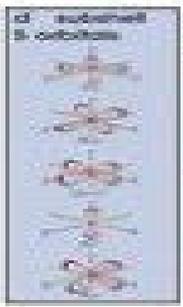
CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1.008** — atomic weight
- H** — atomic symbol
- 1** — atomic number



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For neutral atoms, $Z = +$ the number of electrons. The mass number, A , is the sum of the number of protons and the number of neutrons. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. See also: Atomic Mass, Atomic Weight, Crystal Structure of Solids, Isotopes for a given element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number, n , is the number of nodes in the radial wave function. It is the number of shells. $n = 1, 2, 3, \dots$. The number of orbitals in a shell is n^2 . The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$.

NUCLEAR CHEMISTRY

Nuclear reactions occur when nucleons (protons and neutrons) are added or removed. Nuclear energy is released or absorbed. The mass number, A , is the sum of the number of protons and neutrons. The atomic number, Z , is the number of protons. The number of neutrons is $A - Z$.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the electronic structure of atoms. The number of orbitals in a shell is n^2 . The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$.

n	l	m _l	orbitals
1	0	0	1 orbital
2	0	0	1 orbital
2	1	0, ±1	3 orbitals
3	0	0	1 orbital
3	1	0, ±1	3 orbitals
3	2	0, ±1, ±2	5 orbitals
4	0	0	1 orbital
4	1	0, ±1	3 orbitals
4	2	0, ±1, ±2	5 orbitals
4	3	0, ±1, ±2, ±3	7 orbitals

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Jane Gardner



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