

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
 potassium-40 undergoes beta decay (electron)  
 uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.  
 Write the electron configuration for Cr.  
 Write the shorthand notation for Bi.



### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds:

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



# Chemistry Chapter 10 Section 3 Review Answers

**Kathleen Armour**



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**14 5 reactions of acids and bases chemistry libretexts** - Jan 07 2023

web acids and bases react with metals acids react with most metals to form a salt and hydrogen gas as discussed previously metals that are more active than acids can

**acid and base strength chemistry libretexts** - May 31 2022

web jan 30 2023 acids and bases behave differently in solution based on their strength acid or base strength is a measure of how readily the molecule ionizes in water introduction

**14 3 relative strengths of acids and bases chemistry libretxts** - Apr 10 2023

web jul 20 2023 assess the relative strengths of acids and bases according to their ionization constants rationalize trends in acid base strength in relation to molecular structure

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web reinforcement acids and bases apr 12 2020 acids and bases are ubiquitous in chemistry our understanding of them however is dominated by their behaviour in

**bases and alkalis acids and bases ks3 chemistry revision** - Apr 29 2022

web a base is a substance that can react with acids and neutralise them bases are usually metal oxides such as copper oxide metal hydroxides such as sodium hydroxide or

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web an i acid bpse neutralization reaction is 2 a n is a substance that conductp electricity 3 the chloride ion cl is the l base of hydrochloric acid hcl 4 the formula hro