

Chemistry

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(Chapter 1)(Solutions) (InText Questions)
(Class 12)

Question 1.11:

Calculate the mass of ascorbic acid (Vitamin C, $C_6H_8O_6$) to be dissolved in 75 g of acetic acid to lower its melting point by 1.5°C . $K_f = 3.9 \text{ K kg mol}^{-1}$.

Answer 1.11:

Mass of acetic acid, $w_1 = 75 \text{ g}$

Molar mass of ascorbic acid ($C_6H_8O_6$), $M_2 = 6 \times 12 + 8 \times 1 + 6 \times 16 = 176 \text{ g mol}^{-1}$

Lowering of melting point, $\Delta T_f = 1.5 \text{ K}$

We know that:

$$\Delta T_f = \frac{K_f \times 1000 \times w_2}{M_2 \times w_1}$$

$$\Rightarrow w_2 = \frac{\Delta T_f \times M_2 \times w_1}{K_f \times 1000}$$

$$= \frac{0.5 \times 176 \times 75}{3.9 \times 1000}$$

$$= 5.08 \text{ g (approx)}$$

Hence, 5.08 g of ascorbic acid is needed to be dissolved.

Question 1.12:

Calculate the osmotic pressure in Pascals exerted by a solution prepared by dissolving 1.0 g of polymer of molar mass 185,000 in 450 mL of water at 37°C .

Answer 1.12:

It is given that:

Volume of water, $V = 450 \text{ mL} = 0.45 \text{ L}$

Temperature, $T = (37 + 273) \text{ K} = 310 \text{ K}$

Number of moles of the polymer, $n = \frac{1}{185000} \text{ mol}$

We know that: $\pi = \frac{n}{V} RT$

Osmotic pressure,

$$= \frac{1}{185000} \text{ mol} \times \frac{1}{0.45 \text{ L}} \times 8.314 \times 10^3 \text{ Pa L K}^{-1} \times 310 \text{ K}$$

$$= 30.98 \text{ Pa}$$

$$= 31 \text{ Pa (approximately)}$$

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Chemistry Chapter 12 Solutions Answers

Darrell D. Ebbing



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