

## Chapter Test II

### A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

#### Column I

- \_\_\_\_\_ 1. An oxidation-reduction reaction in which an element is oxidized and another element is reduced.
- \_\_\_\_\_ 2. An equation of a redox reaction that is balanced for both mass and charge.
- \_\_\_\_\_ 3. The reduction of an oxidizing agent.
- \_\_\_\_\_ 4. The ratio of the amount of substance that is oxidized to the amount that is reduced.
- \_\_\_\_\_ 5. An equation that is balanced for both mass and charge.
- \_\_\_\_\_ 6. An equation that is balanced for mass but not for charge.

#### Column II

- a. oxidation
- b. net ionic equation
- c. reduction
- d. half-reaction
- e. balanced equation
- f. oxidation-reduction reaction

### B. Multiple Choice

Choose the best answer for each question.

- \_\_\_\_\_ 1. Which of the following is not a redox reaction?
  - a.  $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$
  - b.  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
  - c.  $\text{H}_2\text{O} + \text{CO}_2 \rightarrow \text{H}_2\text{CO}_3$
  - d.  $\text{Fe} + \text{Cu}^{2+} \rightarrow \text{Fe}^{2+} + \text{Cu}$
- \_\_\_\_\_ 2. Which of the following is not a redox reaction?
  - a.  $\text{Fe} + \text{Cu}^{2+} \rightarrow \text{Fe}^{2+} + \text{Cu}$
  - b.  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
  - c.  $\text{H}_2\text{O} + \text{CO}_2 \rightarrow \text{H}_2\text{CO}_3$
  - d.  $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$
- \_\_\_\_\_ 3. Which of the following is not a redox reaction?
  - a.  $\text{Fe} + \text{Cu}^{2+} \rightarrow \text{Fe}^{2+} + \text{Cu}$
  - b.  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
  - c.  $\text{H}_2\text{O} + \text{CO}_2 \rightarrow \text{H}_2\text{CO}_3$
  - d.  $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$
- \_\_\_\_\_ 4. Which of the following is not a redox reaction?
  - a.  $\text{Fe} + \text{Cu}^{2+} \rightarrow \text{Fe}^{2+} + \text{Cu}$
  - b.  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
  - c.  $\text{H}_2\text{O} + \text{CO}_2 \rightarrow \text{H}_2\text{CO}_3$
  - d.  $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$

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