

Nomenclature Review

Name: _____

1. Give the charge and name for each of the following *oxyanions*.

	Charge	Name
a. NO_3^-	-1	nitrate
b. CO_3^{2-}	-2	carbonate
c. $\text{C}_2\text{H}_3\text{O}_2^-$	-1	acetate
d. ClO_2^-	-1	chlorite
e. PO_3^{2-}	-3	phosphite
f. ClO_4^-	-1	perchlorate
g. SO_4^{2-}	-2	sulfate
h. NH_4^+	+1	ammonium
i. OH^-	-1	hydroxide
j. CN^-	-1	cyanoide
k. BrO_3^-	-1	bromate
l. $\text{C}_2\text{O}_4^{2-}$	-2	oxalate
m. SO_3^{2-}	-2	sulfite
n. MnO_4^-	-1	permanganate

2. If A and B represent compounds of an element bonded to a polyatomic ion, give an example of a valid formula for each combination below.

a. AB	b. AB_2	c. A_2B	d. AB_3	e. A_2B_3
$\text{Al}(\text{NO}_3)_3$	$\text{Ca}(\text{NO}_2)_2$	$(\text{NH}_4)_2\text{S}$	$\text{Al}(\text{ClO}_4)_3$	$\text{Al}_2(\text{SO}_4)_3$

3. What is the oxidation number of:

a. Calcium ion?	+2		
b. Sodium ion?	+1		
c. Sulfur ion?	-2		
d. Phosphorus ion?	-3		
e. Carbon ion within Carbon dioxide molecule?	+4	$x + 2(-2) = 0$ $x = 4+$	CO_2
f. Phosphorus within the Phosphate polyatomic ion?	+5		PO_4^{3-} $x + 4(-2) = -3$ $x = +6$
g. Chromium within the Chromate polyatomic ion?	+6		
h. Carbon within the Oxalate polyatomic ion?	+3		$\text{C}_2\text{O}_4^{2-}$ $2x + 4(-2) = -2$ $x = 3+$
i. Zinc ion?	+2		
j. Cadmium ion?	+2		
k. Sulfur within the Hydrogen Sulfite ion (HSO_3^-)	+4	$1 + x - 6 = -1$ $x = 4+$	

4. Give the most common oxidation number/numbers for each of the following groups or families of elements.

a. Alkali Metals	+1
b. Alkaline Earth	+2
c. Noble Gases	0
d. Halogens	-1
e. Group 13	+3
f. Group 15	-3

5. Give the formula for the following hydrocarbons/alcohols:

a. methane	C_1H_4
b. butene	C_4H_8
c. ethanol	$\text{C}_2\text{H}_5\text{OH}$
d. propyne	C_3H_4

Chemistry Unit 4 Test Review Electron Configuration

Rosina Ehmman



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