

Chemistry – Unit 7 Chemical Reactions

Lesson 1: Atoms

Background

Chemistry before 1900 generally focused on the study of elements. The scientific approach to study of the atom is recent.

Historical View of Atomic Model of Matter

The atomic theory of matter postulates that all matter consists of very small particles which are constantly moving in a continuous motion. The discovery of the particulate nature of matter led to the development of the atomic theory that deals with the stability of atoms.

1. Conservation of Mass

A principle stating that matter neither is created nor destroyed.

also in the real world

2. Chemical Formula

A set of chemical symbols representing atoms present in a compound and

their relative proportions, which define the structure of the compound.

3. Molar Mass and Molar Ratio

Relative number of molecules or atoms in a compound indicates the

number of atoms of each element present in any molecule or formula unit.

also

Chemical formula gives us the formula of a compound in writing.

A molecule is formed that characterizes a compound.

also

4. Pressure

1. The gases were studied by the scientists like Robert Boyle, Jacques Charles, Joseph Gay-Lussac. They discovered the relation between the pressure and volume of a gas. Boyle's law was first stated by Robert Boyle in 1662. It states that the pressure of a gas is inversely proportional to its volume.

2. The gas particles themselves do not exert any pressure. However, they exert a pressure on the walls of the container.

3. Chemists discovered that all gaseous molecules were made of either two or more than two atoms. It was found by Avogadro's law that the volume of a gas is directly proportional to the number of molecules.

4. Chemists have found gaseous molecules are made of either two or more than two atoms. It was found that the volume of a gas is directly proportional to the number of molecules.

Chemistry Unit 7 Rearranging Atoms Answers

JL Elias



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