









VSEPR Theory (Molecular Shapes)

A = the central atom, X = an atom bonded to A, E = a lone pair on A

Note: There are lone pairs on X or other atoms, but we don't care. We are interested in only the electron densities or domains around atom A.

Total Domains	Generic Formula	Picture	Bonded Atoms	Lone Pairs	Molecular Shape	Electron Geometry	Example	Hybridization	Bond Angles
1	AX	A—X	1	0	Linear	Linear	H ₂	s	180
2	AX ₂	X—A—X	2	0	Linear	Linear	CO ₂	sp	180
	AXE	 A—X	1	1	Linear	Linear	CN		
3	AX ₃		3	0	Trigonal planar	Trigonal planar	AlBr ₃	sp ²	120
	AX ₂ E		2	1	Bent	Trigonal planar	SnCl ₂		
	AXE ₂		1	2	Linear	Trigonal planar	Cl ₂		
4	AX ₄		4	0	Tetrahedral	Tetrahedral	SiCl ₄	sp ³	109.5
	AX ₃ E		3	1	Trigonal pyramidal	Tetrahedral	NH ₃		
	AX ₂ E ₂		2	2	Bent	Tetrahedral	H ₂ O		
	AXE ₃		1	3	Linear	Tetrahedral	Cl ₄		

Chemistry Vsepr Worksheet Answers

L Reisser



Chemistry Vsepr Worksheet Answers:

Chemistry Vsepr Worksheet Answers Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chemistry Vsepr Worksheet Answers**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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