

REACTIVITY TRENDS IN THE PERIODIC TABLE

Reactivity Trends in the Periodic Table

Reading Passage

Read the passage carefully before answering the questions that follow.

The periodic table organizes elements based on their atomic number and properties, revealing patterns known as periodic trends. One of these trends is reactivity, which describes how likely an element is to participate in chemical reactions. Understanding these trends helps scientists predict how elements will behave in various chemical contexts.

Metals, such as sodium, calcium, and potassium, are highly reactive because they have a single valence electron that they readily lose in chemical reactions. This tendency is due to their large atomic size and low ionization energy, making it energetically favorable for them to react with other substances, especially nonmetals. Conversely, noble gases, like helium, neon, and argon, are very inert because they have full valence electron shells. This stable electronic configuration makes them unlikely to form bonds or participate in reactions under normal conditions.

Across periods in the periodic table, reactivity generally decreases as atomic number increases and valence electrons are held more tightly. In addition, atomic size influences reactivity; larger atoms tend to be more reactive if they have fewer valence electrons in their outermost shell. This electronic configuration directly impacts their ability to attract or give up electrons during chemical reactions. The periodic table also illustrates trends in reactivity for nonmetals, which generally increase as atomic number increases and atomic size decreases. These trends are closely related to electronegativity, which measures an atom's ability to attract electrons. Understanding these patterns, scientists can better predict and control chemical reactions, including developments in material science, industrial chemistry, and environmental science.

True or False: Reactivity in the Periodic Table

Circle whether each statement is True or False.

- Metals are highly reactive because they have single valence electrons that are difficult to lose.
- Noble gases are very reactive because they have full valence electron shells.
- Reactivity for elements across a period which generally decreases as the number of elements that increase.
- Metals and their valence electrons tend to be held more tightly than their outermost electrons.

Periodic Trends Reactivity Lab Answer Key

Victor M. Corman



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