

Chapter 5 Chapter Review Question Answers

- Describe the scientists' contributions:
 - Cannizzaro—developed a standard method for measuring atomic masses, which allowed chemists to search for periodic trends among elements.
 - Mendeleev—organized elements according to increasing atomic mass and noticed that similar properties appeared periodically.
 - Moseley—discovered that atomic number, not atomic mass, should be the basis of organizing elements on the periodic table.
- The physical and chemical properties of the elements are periodic functions of their atomic numbers.
- Groups of elements exhibit similar chemical and physical properties and behaviors.
- Electron configurations:
 - Generally, the configurations of the outermost electrons shells of elements within the same group are the same. There are a number of exceptions to this rule, however, among the transition elements.
 - Their outer shells are completely filled.
- The length of a period is determined by the total number of electrons that can fill the outer sublevels of the elements of that period.
- An element's period corresponds to its highest occupied main energy level. (s & p)
- Blocks & Groups
 - The type of sublevel being filled in successive elements of that block.
 - s-block: Groups 1 & 2; p-block: Groups 13–18; d-block: Groups 3–12; f-block: No group #s, as they are between groups 3 & 4, but period 6 elements are in the Lanthanide series, and period 7 elements are in the Actinide series.
- Alkali Metals
 - The Group 1 elements, except Hydrogen.
 - Alkali metals are extremely reactive, they react violently with water, they have low densities, they are soft enough to be cut with a knife, and freshly cut alkali metals are very shiny silver colored metals.
- Alkaline Metals
 - The Group 2 elements.
 - They are slightly less reactive than Group 1 elements; they are harder, duller, denser than Group 1 elements. They also have slightly higher melting points than the Group 1 elements.
- Group Configurations:
 - Transition Metals (d-block)
 - Group 3: $ns^2 (n-1)d^1$
 - Group 4: $ns^2 (n-1)d^2$
 - Group 5: $ns^2 (n-1)d^3$
 - Group 6: $ns^2 (n-1)d^4$ {actually is $ns^1 (n-1)d^5$ for Cr & Mo}
 - Group 7: $ns^2 (n-1)d^5$
 - Group 8: $ns^2 (n-1)d^6$
 - Group 9: $ns^2 (n-1)d^7$
 - Group 10: $ns^2 (n-1)d^8$
 - Group 11: $ns^2 (n-1)d^9$ {actually is $ns^1 (n-1)d^{10}$ for all group 11 elements}
 - Group 12: $ns^2 (n-1)d^{10}$
 - The group number is the sum of the s & d electrons.
- The transition elements.
- p-block
 - The p-block consists of nonmetals at the right, metalloids in the middle along the staircase, and metals at the left.
 - The p-block metals are generally harder and denser than the s-block metals, but softer and less dense than the d-block metals.
- Halogens
 - The Group 17 elements.

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Charles H. Corwin



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Solutions to Exercises Roxy Wilson, 2003 Full solutions to all end of chapter exercises in the text are provided With an instructor's permission this manual may be made available to students

Academic Writing for International Students of Science Jane Bottomley, 2014-07-25 Academic Writing for International Students of Science will help international students to develop their command of academic scientific writing in English It guides students through the writing process itself and will help them to produce clear well written and well organised essays and reports The book covers a range of issues such as how to explain complex ideas clearly and concisely how to develop a coherent argument and how to avoid plagiarism by making effective reference to sources Through detailed analysis of authentic scientific texts the book will enhance students understanding of the nature of academic scientific writing This will enable them to understand how language and discourse function in a real scientific context The texts serve as models of good writing and are followed by practice activities which will help students to develop their own writing skills Key topics include the writing process academic scientific style sentence structure paragraph development referring to sources coherence argument and critical thinking academic and scientific conventions This book will be an invaluable companion to those studying for a science or technology degree in an English speaking institution Informative study boxes model answers and a clear comprehensive answer key mean that the book can be used for self study or with guidance in the classroom

The School Science Review, 2002

Introductory Chemistry Charles H. Corwin, 2005 For one semester courses in Basic Chemistry Introduction to Chemistry and Preparatory Chemistry and the first term of Allied Health Chemistry This text is carefully crafted to help students learn chemical skills and concepts more effectively Corwin covers math and problem solving early in the text he builds student confidence and skills through innovative problem solving pedagogy and technology formulated to meet student needs

Chemists' Guide to Effective Teaching Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 Part of the Prentice Hall Series in Educational Innovation for Chemistry this unique book is a collection of information examples and references on learning theory teaching methods and pedagogical issues related to teaching chemistry to college students In the last several years there has been considerable activity and research in chemical education and the materials in this book integrate the latest developments in chemistry Each chapter is written by a chemist who has some expertise in the specific technique discussed has done some research on the technique and has applied the technique in a chemistry course

Chemistry - The Central Science James C. Hill, Bruce Edward Bursten, 2006 Chemistry The Central Science is the most trusted book on the market its scientific accuracy clarity innovative pedagogy functional problem solving and visuals set this book apart Brown LeMay and Bursten teach students the concepts and skills they need without overcomplicating the subject A comprehensive media package that works in tandem with the text helps students practice and learn while providing instructors the tools they need to succeed Publisher's description

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Chemometrics for Analytical Chemistry James N. Miller, Jane Charlotte Miller, 2000 This popular textbook gives a clear and lucid account of the underlying principles of statistical methods The fourth edition has been revised and updated to reflect the growing popularity of statistics and chemometric methods and new approaches in optimization and experimental design The authors have also addressed the quality of analytical chemistry data and experimental results an area of increasing concern to chemists testing the safety of food and medicines This book will suit undergraduate M Sc and graduate courses in Analytical Chemistry and related topics and will also be valuable for researchers and chemists working in analytical chemistry everywhere

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