

Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

Chemical Bonding:

- **Four types of bonds:**
 - *Nonpolar Covalent Bond* (ΔEN (difference in electronegativity) = < 0.3)
 - **EQUAL SHARING** of electrons between the 2 bonding nonmetal atoms
 - *Polar Covalent Bond* ($\Delta EN = 0.4-1.7$)
 - **UNEQUAL SHARING** of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
 - *Ionic Bond* ($\Delta EN = > 1.7$)
 - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely **GIVING UP/TRANSFERRING ELECTRONS** to the atom with the greater EN value
 - Ionic bond results due to the attraction between the **POSITIVE** cation (formed by losing electrons) and the **NEGATIVE** anion (formed by gaining electrons)
 - *Metallic Bond*
 - Bonding between two metallic atoms which occurs due to the attraction between a “sea of electrons” and the nuclei of bonded metal atoms
- * Must memorize/know the difference in electronegativity ranges for each type of bond
 - ** Note: you should still know the definition of electronegativity (*EN*)
- **Properties of the Four Types of Chemical Bonds**
 - See the table at the bottom of Packet pg.5
 - You **MUST** be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
 - **Why** → because all atoms want to be at their **LOWEST POTENTIAL ENERGY**; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
 - **How** → **ATTRACTIVE** and **REPULSIVE** forces
 - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
 - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
 - **Bond Length**
 - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
 - **Bond Energy**
 - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

Chapter 6 Study Guide Chemistry

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Chapter 6 Study Guide Chemistry:

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Study Guide [to Accompany] General Chemistry Dixie Goss, 2002

Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988

Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter, 1993

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the authors Also of benefit is the just in time presentation of key content only providing details once they are needed While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons The whole is backed by a range of supplements including a new illustration program a tutorial CD interactive learningware an extensive Web CT component an instructor s resource CD and a solution CD

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Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry Iris Stovall,Donald J. DeCoste,2000

Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual T. W. Graham Solomons,1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry McMurry,Susan McMurry,1990

Robinson Chemistry Study Guide Robinson,1992

Study Guide, Student Edition, for Use with Glencoe Life Science McGraw Hill,1998-05

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veterans day 2014 calendar date - Feb 19 2022

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