

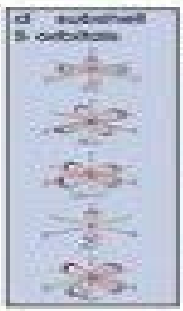
CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1 - Atomic Number
- H - Atomic Symbol
- 1.008 - Atomic Weight



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For neutral atoms, $Z = +$ the number of electrons. The mass number, A , is the sum of the number of protons and the number of neutrons. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. The atomic weight is the weighted average of isotopes for a given element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number, n , is the number of shells. It is the number of nodes in the radial wave function. It is the number of nodes in the radial wave function. It is the number of nodes in the radial wave function.

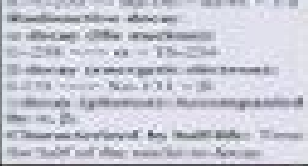
NUCLEAR CHEMISTRY

Nuclear reactions occur when nucleons (protons and neutrons) are added or removed. Nuclear energy is released or absorbed. Nuclear reactions are characterized by ΔE and Δm .

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the many-electron atoms. The Aufbau principle is used to determine the ground state configuration of an atom.

n	l	m _l	m _s	Orbital
1	0	0	±1/2	1s
2	0	0	±1/2	2s
2	1	0, ±1	±1/2	2p
3	0	0	±1/2	3s
3	1	0, ±1	±1/2	3p
3	2	0, ±1, ±2	±1/2	3d



Block	Group	Element	Symbol	Atomic Number		
s-block	1	H	H	1		
		Li	Li	3		
	2	Be	Be	4		
		Mg	Mg	12		
	p-block	13	B	B	5	
			Al	Al	13	
		14	C	C	6	
			Si	Si	14	
		15	N	N	7	
			P	P	15	
		16	O	O	8	
			S	S	16	
		17	F	F	9	
			Cl	Cl	17	
		18	Ne	Ne	10	
			Ar	Ar	18	
		d-block	8	Sc	Sc	21
				Ti	Ti	22
V	V			23		
Cr	Cr			24		
Mn	Mn			25		
Fe	Fe			26		
Co	Co			27		
Ni	Ni			28		
Cu	Cu			29		
Zn	Zn			30		
f-block	7	La	La	57		
		Ce	Ce	58		
		Pr	Pr	59		
		Nd	Nd	60		
		Pm	Pm	61		
		Sm	Sm	62		
		Eu	Eu	63		
		Gd	Gd	64		
		Tb	Tb	65		
		Dy	Dy	66		
		Ho	Ho	67		
		Er	Er	68		
		Tm	Tm	69		
		Yb	Yb	70		
Lu	Lu	71				

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Michael Rajnik



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laminated guide is intended to serve as a reminder and quick study review for all aspects of the chemistry lab from safety to chemicals to instruments and safety again first last and always Designed to find answers fast and to be spill proof it s perfect for high school and early college students studying chemistry as well as anyone interested in learning how scientists safely study the molecules that make up everything around us 6 page laminated guide includes Lab Safety Training General Lab Guidelines Working With Chemicals Exposure to Chemicals First Aid Know Your Lab Reagents Chemical Spills Waste Management Useful Chemical Information Periodic Table Included Lab Must Knows Safe Use of Lab Equipment Data Manipulation Preparing a Solution **Inorganic Chemistry** Mark Jackson,2010-05-31 Our 3 panel 6 page guide is jam packed with chemistry information not covered in organic chemistry and biochemistry Perfect for science students who want to focus on the aspects of chemistry that go beyond what is covered in material that deals with organic topics this guide covers the key concepts principles figures and formulas that inorganic chemistry students will need to know in order to succeed Tables images and graphic elements further enhance the text *Chemistry 2* BarCharts, Inc,Jane P. Gardner,2016

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