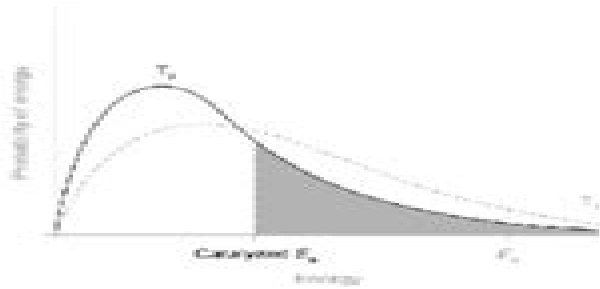


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2.	f	<p>a. <math>Zn^{2+}</math> ions in full shell  <b>OR</b>                      dentic number is zero since no lone pairs is there</p>	<p>Do not accept <math>Zn</math> in any transition metal.                      Do not accept structures for <math>Zn^{2+}</math> ions.</p>	1
3.	b	<p>ligands donate pairs of electrons to metal ions  <b>OR</b>                      form coordinate covalent bonds</p> <p>ligands are Lewis bases  <b>AND</b>                      metal ions are Lewis acids</p>		1
1.	d	<p>Electrostatic attraction                      Between oppositely charged ions/between <math>Fe^{2+}</math> and <math>SO_4^{2-}</math></p>		2
4.	a	 <p>same higher <b>AND</b> to left of <math>T_1</math> if reactant <math>E_a</math> reduced <b>AND</b> to the left of <math>E_a</math> of curve <math>T_1</math> if</p>	<p>Do not promote curve into a table, and justify clearly (height and width, or stating it is after it is)</p> <p>Do not award AP if curve shown shows significantly more than molecular/molar/molar area under curve than curve 1.</p> <p>Accept <math>E_a</math> closer to <math>T_1</math> instead of curve shown or closer to left of modified <math>T_1</math></p>	2
1.	b	<p>to same <b>AND</b> same <math>T_1</math></p>		1

3. Equilibrium favours backward direction/shifts to LHS **AND**

As increase in pressure favours side with less moles of gas.

[1]

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