

## 9-2 Review and Reinforcement

### Chemical Equations

Refer to the unbalanced chemical equation below to answer each of the following questions. Write your answers in the space provided.



- Write the names of the reactants in this equation.  
Aluminum Nitrate + Iron (II) chloride
- Write the names of the products in this equation.  
Iron (II) Nitrate + Aluminum chloride
- What would the coefficient of  $\text{Al(NO}_3)_3$  be in the balanced equation? Two
- What would the coefficient of  $\text{Fe(NO}_3)_2$  be in the balanced equation? Three
- What would the coefficient of  $\text{AlCl}_3$  be in the balanced equation? Two
- How many oxygen atoms would there be on each side of the balanced equation?  
18

Answer each of the following questions in the space provided.

- What should you do if an equation you have written cannot be balanced?  
make sure reactants + products are written correctly, try again, ask for help
- How is balancing a chemical reaction related to the law of conservation of matter?  
mass of reactants must = mass of products
- Describe the correct way to balance an equation.  
add only coefficients never change formulas or subscripts
- How does a chemical equation relate the macroscopic, observable reaction to the interactions occurring at the microscopic level?  
Same

# 3 1 Review Reinforcement Answer Key

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**The American Community Survey**, 2004

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