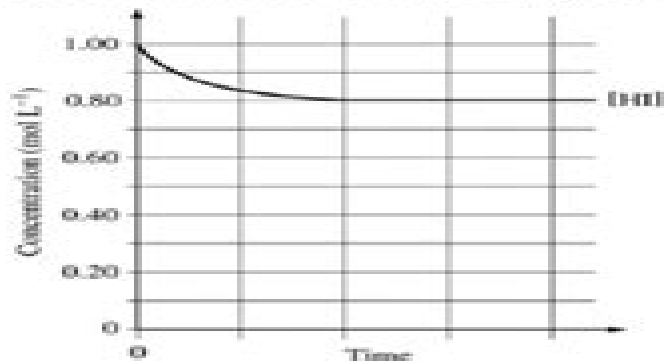


## AP Chemistry Unit 7 General Equilibrium Free Response Practice

### Question 1.



After a 1.0 mole sample of  $\text{HI}(\text{g})$  is placed into an evacuated 1.0 L container at 700. K, the reaction represented above occurs. The concentration of  $\text{HI}(\text{g})$  as a function of time is shown below.



(b) What is  $[\text{HI}]$  at equilibrium?

(c) Determine the equilibrium concentrations of  $\text{H}_2(\text{g})$  and  $\text{I}_2(\text{g})$ .

(d) On the graph above, make a sketch that shows how the concentration of  $\text{H}_2(\text{g})$  changes as a function of time.

(e) Calculate the value of the following equilibrium constants at 700. K.

(i)  $K_c$

(ii)  $K_p$

(f) At 1,000 K, the value of  $K_c$  for the reaction is  $2.6 \times 10^{-2}$ . In an experiment, 0.75 mole of  $\text{HI}(\text{g})$ , 0.10 mole of  $\text{H}_2(\text{g})$ , and 0.50 mole of  $\text{I}_2(\text{g})$  are placed in a 1.0 L container and allowed to reach equilibrium at 1,000 K. Determine whether the equilibrium concentration of  $\text{HI}(\text{g})$  will be greater than, equal to, or less than the initial concentration of  $\text{HI}(\text{g})$ . Justify your answer.

# 7 Ap Chemistry Response Answers

**S Baum**



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