

4. A student adds 15.0 mL of water to a 25.0 mL sample of 0.80 M Li_3PO_4 solution.

$$FC = IC \times \frac{V}{FV}$$

- a) Calculate the final $[\text{Li}^+]$ (2 marks)

$$\text{Final } [\text{Li}_3\text{PO}_4] = 0.80 \text{ M} \times \frac{25.0 \text{ mL}}{40.0 \text{ mL}} = 0.50 \text{ M}$$



$$\text{final } [\text{Li}^+] = 1.5 \text{ M}$$

- b) Calculate the final $[\text{PO}_4^{3-}]$ (1 mark)



$$\text{final } [\text{PO}_4^{3-}] = 0.50 \text{ M}$$

5. Calculate the mass of solid K_2CO_3 which must be added to 500.0 mL of water in order to make a solution in which the $[\text{K}^+]$ is 3.0 M. (3 marks)



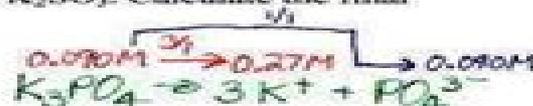
$$\text{mol} = \text{M} \times \text{L} = 1.50 \text{ M} \times 0.5000 \text{ L} = 0.750 \text{ mol } \text{K}_2\text{CO}_3$$

$$\text{mass} = 0.750 \text{ mol } \text{K}_2\text{CO}_3 \times \frac{138.2 \text{ g } \text{K}_2\text{CO}_3}{1 \text{ mol } \text{K}_2\text{CO}_3} = 103.65 \text{ g}$$

$$\text{mass } \text{K}_2\text{CO}_3 = 104 \text{ g}$$

6. 60.0 mL of 0.45 M K_3PO_4 is mixed with 240.0 mL of 0.20 M K_2SO_4 . Calculate the final concentration of all three ions in the solution. (3 marks)

$$\text{final } [\text{K}_3\text{PO}_4] = 0.45 \text{ M} \times \frac{60.0 \text{ mL}}{300.0 \text{ mL}} = 0.090 \text{ M}$$



$$\text{final } [\text{K}_2\text{SO}_4] = 0.20 \text{ M} \times \frac{240.0 \text{ mL}}{300.0 \text{ mL}} = 0.16 \text{ M}$$



$$\text{final } [\text{K}^+] = 0.27 \text{ M} + 0.32 \text{ M} = 0.59 \text{ M}$$

$$\text{Answer: } [\text{PO}_4^{3-}] = 0.090 \text{ M} \quad [\text{SO}_4^{2-}] = 0.16 \text{ M} \quad [\text{K}^+] = 0.59 \text{ M}$$

Acid Base Titration Lab Answer Key

Richard W. Ramette



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