

Acid-Base Titration Practice Problems

- 1) Calculate the pH during the titration of 40.00 mL of 0.0500 M butanoic acid, $\text{CH}_3\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$, with 0.1000 M NaOH solution after the following additions of the base: a) 0.00 mL; b) 15.00 mL; c) 19.00 mL; d) 20.00 mL; e) 25.00 mL. The K_a for butanoic acid is 1.54×10^{-5} .
- 2) Calculate the pH during the titration of 25.00 mL of 0.200 M KOH with 0.100 M HClO_4 after the following additions of acid: a) 0.00 mL; b) 25.00 mL; c) 39.00 mL; d) 50.00 mL; e) 60.00 mL.
- 3) Calculate the pH when the following quantities of 0.0500 M $\text{Ba}(\text{OH})_2$ solution have been added to 25.00 mL of 0.0400 M HNO_3 solution: a) 0.00 mL; b) 5.00 mL; c) 10.00 mL; d) 20.00 mL.

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