

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the vocabulary list.

- | | |
|-------------------------|---------------------------|
| _____ mole | _____ percent composition |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ molar mass |
| _____ molar mass | _____ molar mass |
10. _____ is a description of the relative amounts of each element in a compound. _____ percent composition
11. _____ is the lowest whole number ratio of the atoms of the elements in a compound. _____ empirical formula
12. _____ is the mass of one mole of any gas at STP. _____ molar mass
13. _____ is the mass of one mole of a compound. _____ molar mass
14. _____ is the amount representing $n \times 6.02 \times 10^{23}$ representative particles of a substance. _____ mole
15. _____ is 6.02×10^{23} particles. _____ Avogadro's number
16. _____ is the mass of one mole of a substance at STP. _____ molar mass
17. _____ is the mass of one mole of a gas at STP. _____ molar mass
18. _____ is the mass of one mole of a gas at STP. _____ molar mass

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Complete the following practice problems using your knowledge of the terms and your understanding of the concepts involved with each of the terms. Write your answers in the space provided or on a separate sheet of paper.

1. _____ is the amount representing $n \times 6.02 \times 10^{23}$ representative particles of a substance. _____ mole

2. _____ is the mass of one mole of a compound. _____ molar mass

3. _____ is the mass of one mole of a compound. _____ molar mass

4. _____ is the mass of one mole of a compound. _____ molar mass

5. _____ is the mass of one mole of a compound. _____ molar mass

Practice Problems

1. _____ is the mass of one mole of a compound. _____ molar mass
2. _____ is the mass of one mole of a compound. _____ molar mass
3. _____ is the mass of one mole of a compound. _____ molar mass
4. _____ is the mass of one mole of a compound. _____ molar mass

Answers

Chapter 10 Chemical Quantities D Practice Answers

JR Anderson



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Introduction to General, Organic, and Biochemistry Morris Hein, 1997. The authors recognize that both science and mathematics may be daunting subjects for many students taking this course. With this in mind they have anticipated where students might stumble and have paced and organized this text to help them through. Their goal is to make the material interesting and relevant so students understand the basic chemical principles related to their career. The authors emphasize problem solving and provide a range of practice exercises. As in previous editions the text first presents the basic concepts of general chemistry and then moves into organic and biochemistry. In this edition the first two sections have been revised primarily to improve explanations and include new pedagogical features. The biochemistry portion has been thoroughly updated to include coverage of many recent developments and emerging technologies in the field.

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