

Chemistry Student Edition - Basic Answer Key

Chapter 11: Chemical Reactions

Chemical Equations

Questions

1. Early Chinese alchemists discovered an early form of gunpowder. What was the composition of this substance?
2. What later developments were made to the gunpowder recipe that improved its pyrotechnic properties?
3. Make an argument for why the burning of a candle is consistent with the law of conservation of matter/mass.
4. Think of an experiment that you could conduct to demonstrate that mass is conserved for a given chemical change.
5. Balance the following chemical equations:
 - a. $C + O_2 \rightarrow CO$
 - b. $CO + O_2 \rightarrow CO_2$
 - c. $H_2 + Br_2 \rightarrow HBr$
 - d. $K + H_2O \rightarrow KOH + H_2$
 - e. $O_3 \rightarrow O_2$
 - f. $N_2 + H_2 \rightarrow NH_3$
 - g. $Zn + AgCl \rightarrow ZnCl_2 + Ag$
 - h. $Cl_2 + NaI \rightarrow NaCl + I_2$
 - i. $P_4O_{10} + H_2O \rightarrow H_3PO_4$
 - j. $Be_2C + H_2O \rightarrow Be(OH)_2 + CH_4$
 - k. $S + HNO_3 \rightarrow H_2SO_4 + NO_2 + H_2O$
 - l. $NH_3 + CuO \rightarrow Cu + N_2 + H_2O$
 - m. $HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$

Answers

1. Early gunpowder was comprised of saltpeter (impure potassium nitrate), sulfur, and charcoal (carbon).
2.
 - a. Purification of potassium nitrate
 - b. Using potassium chlorate instead of potassium nitrate
 - c. Further refinement of the recipe to find the proportions of reactants that yielded the most energetic reaction
3. The burning of a candle obeys the law of conservation of mass because, even though the mass of the candle will decrease as it burns, the mass is not lost. It is transferred into carbon dioxide and water that enter the surrounding air.
4. The reaction between iron and oxygen to produce rust could be studied as an example of conservation of mass. Steel wool could be placed into a covered beaker containing water. The reaction would proceed over time and rust would begin to form. The beaker could be weighed before and after the change. The mass would be the same, even though a new

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Terry C. Jones



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