

Chapter 12

SECTION 12.1

12.1.1 Boyle's Law: Pressure and Volume of a Gas at Constant Temperature

1. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume at constant temperature. Boyle's Law can be written as $PV = k$, where P is pressure, V is volume, and k is a constant.

Example 1: A gas occupies a volume of 100 cm³ at a pressure of 100 kPa.

Question: What is the volume of the gas if the pressure is increased to 200 kPa?

Answer: Boyle's Law states that $PV = k$. So, $100 \times 100 = k$ and $200 \times V = k$.

Therefore, $100 \times 100 = 200 \times V$. Solving for V , we get $V = 50$ cm³.

2. Boyle's Law can be used to determine the volume of a gas at a different pressure if the temperature and the mass of the gas are constant. Boyle's Law can be written as $P_1V_1 = P_2V_2$, where P_1 and V_1 are the initial pressure and volume, and P_2 and V_2 are the final pressure and volume. Boyle's Law can be used to determine the volume of a gas at a different pressure if the temperature and the mass of the gas are constant.

Example 2: A gas occupies a volume of 100 cm³ at a pressure of 100 kPa. What is the volume of the gas if the pressure is increased to 200 kPa?

Answer: Boyle's Law states that $P_1V_1 = P_2V_2$. So, $100 \times 100 = 200 \times V_2$. Solving for V_2 , we get $V_2 = 50$ cm³.

3. Boyle's Law can be used to determine the volume of a gas at a different pressure if the temperature and the mass of the gas are constant. Boyle's Law can be written as $P_1V_1 = P_2V_2$, where P_1 and V_1 are the initial pressure and volume, and P_2 and V_2 are the final pressure and volume. Boyle's Law can be used to determine the volume of a gas at a different pressure if the temperature and the mass of the gas are constant.

Example 3: A gas occupies a volume of 100 cm³ at a pressure of 100 kPa. What is the volume of the gas if the pressure is increased to 200 kPa?

Answer: Boyle's Law states that $P_1V_1 = P_2V_2$. So, $100 \times 100 = 200 \times V_2$. Solving for V_2 , we get $V_2 = 50$ cm³.

Chapter 11 Review Molecular Composition Of Gases

Section 2 Answers

JG Myers



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