

Exercises

14.1-14.3 Temperature, internal energy and states of matter

1. (a) Molecules (of the liquid) are in random / haphazard motion (AM)
Molecules (of liquid) are smaller than pollen grains.
(b) Increase the temperature (of the liquid)

2. The sum of (the random distribution of) the KE and PE of (the) molecules

3. No change in KE
Because temperature is constant (during melting)
PE of (the molecules) increases (during melting).
The internal energy increases.

14.4 Specific Heat Capacity

1. 35 000 J

2. 21 000 J/g/°C

3. 0.375 kg

4. 60 s

5. a. Electrical energy supplied = Heat taken in by water
 $2.5^2 \times 20 = 25 \times 20^2$
 $2300 \times 2 = 2.5 \times 4200 \times 60$
 $t = 240 \text{ s [1]}$

b. The kinetic energy of the particle is increasing. [1]

6. $2 = 3 \times 60 = 180 \text{ s [1]}$

Chapter 12 Assessment Thermal Energy Answers

JR Anderson



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