

Chapter 12

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- _____ 1. The first step in most stoichiometry problems is to _____.
- | | |
|---|--|
| a. add the coefficients of the reagents | c. convert given quantities to volumes |
| b. convert given quantities to moles | d. convert given quantities to masses |
- _____ 2. In the reaction $2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g)$, what is the ratio of moles of oxygen used to moles of CO_2 produced?
- | | |
|--------|--------|
| a. 1:1 | c. 1:2 |
| b. 2:1 | d. 2:2 |
- _____ 3. How many moles of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, can be "burned" biologically when 10.0 mol of oxygen is available?
- $$\text{C}_6\text{H}_{12}\text{O}_6(l) + 6\text{O}_2(g) \rightarrow 6\text{CO}_2(g) + 6\text{H}_2\text{O}(l)$$
- | | |
|--------------|-------------|
| a. 0.500 mol | c. 33.3 mol |
| b. 1.67 mol | d. 60.0 mol |
- _____ 4. At STP, how many liters of oxygen are required to react completely with 3.0 liters of hydrogen to form water?
- $$2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g)$$
- | | |
|----------|----------|
| a. 1.8 L | c. 2.0 L |
| b. 3.0 L | d. 2.4 L |
- _____ 5. The equation below shows the decomposition of lead nitrate. How many grams of oxygen are produced when 11.5 g NO_2 is formed?
- $$2\text{Pb}(\text{NO}_3)_2(s) \rightarrow 2\text{PbO}(s) + 4\text{NO}_2(g) + \text{O}_2(g)$$
- | | |
|-----------|-----------|
| a. 1.00 g | c. 2.88 g |
| b. 2.00 g | d. 32.0 g |
- _____ 6. Iron(III) oxide is formed when iron combines with oxygen in the air. How many grams of Fe_2O_3 are formed when 16.7 g of Fe reacts completely with oxygen?
- $$4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$$
- | | |
|-----------|-----------|
| a. 32.0 g | c. 47.8 g |
| b. 23.9 g | d. 85.6 g |
- _____ 7. When glucose is consumed, it reacts with oxygen in the body to produce carbon dioxide, water, and energy. How many grams of carbon dioxide would be produced if 45 g of $\text{C}_6\text{H}_{12}\text{O}_6$ completely reacted with oxygen?
- | | |
|----------|---------|
| a. 1.5 g | c. 11 g |
| b. 1.8 g | d. 66 g |
- _____ 8. Aluminum reacts with sulfuric acid to produce aluminum sulfate and hydrogen gas. How many grams of aluminum sulfate would be formed if 250 g H_2SO_4 completely reacted with aluminum?
- $$2\text{Al}(s) + 3\text{H}_2\text{SO}_4(aq) \rightarrow \text{Al}_2(\text{SO}_4)_3(aq) + 3\text{H}_2(g)$$
- | | |
|-----------|----------|
| a. 0.85 g | c. 450 g |
| b. 290 g | d. 870 g |

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