



Carbon tetrachloride ( $\text{CCl}_4$ )

$\mu = 0\text{D}$

$\text{CCl}_4$  is a symmetrical molecule. Therefore, the dipole moments of all four C–Cl bonds cancel each other. Hence, its resultant dipole moment is zero.

As shown in the above figure, in  $\text{CHCl}_3$ , the resultant of dipole moments of two C–Cl bonds is opposed by the resultant of dipole moments of one C–H bond and one C–Cl bond. Since the resultant of one C–H bond and one C–Cl bond dipole moments is smaller than two C–Cl bonds, the opposition is to a small extent. As a result,  $\text{CHCl}_3$  has a small dipole moment of 1.08 D.

On the other hand, in case of  $\text{CH}_2\text{Cl}_2$ , the resultant of the dipole moments of two C–Cl bonds is strengthened by the resultant of the dipole moments of two C–H bonds. As a result,  $\text{CH}_2\text{Cl}_2$  has a higher dipole moment of 1.60 D than  $\text{CHCl}_3$  i.e.,  $\text{CH}_2\text{Cl}_2$  has the highest dipole moment.

Hence, the given compounds can be arranged in the increasing order of their dipole moments as:



Question 10.5:

A hydrocarbon  $\text{C}_5\text{H}_{10}$  does not react with chlorine in dark but gives a single monochloro compound  $\text{C}_5\text{H}_9\text{Cl}$  in bright sunlight. Identify the hydrocarbon.

Answer

A hydrocarbon with the molecular formula,  $\text{C}_5\text{H}_{10}$ , belongs to the group with a general molecular formula  $\text{C}_n\text{H}_{2n}$ . Therefore, it may either be an alkene or a cycloalkane.

Since hydrocarbon does not react with chlorine in the dark, it cannot be an alkene. Thus, it should be a cycloalkane.

Further, the hydrocarbon gives a single monochloro compound,  $\text{C}_5\text{H}_9\text{Cl}$  by reacting with chlorine in bright sunlight. Since a single monochloro compound is formed, the hydrocarbon must contain H-atoms that are all equivalent. Also, as all H-atoms of a cycloalkane are equivalent, the hydrocarbon must be a cycloalkane. Hence, the said compound is cyclopentane.



Cyclopentane ( $\text{C}_5\text{H}_{10}$ )

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