

## Chapter 10 - Chemical Quantities

### 10.1 The Mole: A Measurement of Matter:

#### Mole:

- A counting unit
- $6.02 \times 10^{23}$  representative particles (atom, molecule, formula unit, ion) of a substance = Avogadro's Number

#### Converting # of particles to Moles:

$$\text{Moles} = \cancel{\text{Representative particles}} \left( \frac{1 \text{ mole}}{6.02 \times 10^{23} \cancel{\text{ representative particles}}} \right)$$

#### Converting Moles to # of particles:

$$\text{Representative particles} = \cancel{\text{moles}} \left( \frac{6.02 \times 10^{23} \cancel{\text{ representative particles}}}{1 \cancel{\text{ mole}}} \right)$$

#### Molar Mass:

- Molar mass of an element = atomic mass (g)  
Ex. 24.3 g = molar mass of Mg
- Molar mass = g/mol
- Molar mass of any element contains 1 mol or  $6.02 \times 10^{23}$  atoms of that element
- Molar mass of a compound = sum of the atomic masses of each element contained in 1 mole of the compound.

### 10.2 Mole-Mass and Mole-Volume Relationships:

#### Mole-Mass Relationship:

- Use molar mass (g/mol) of an element or compound to convert b/w the mass of a substance and the moles of a substance

$$\text{Mass (grams)} = \cancel{\# \text{ of moles}} \left( \frac{\cancel{\text{mass (grams)}}}{1 \cancel{\text{ mole}}} \right)$$

$$\# \text{ of Moles} = \cancel{\text{Mass (grams)}} \left( \frac{1 \cancel{\text{ mole}}}{\cancel{\text{mass (grams)}}} \right)$$

# Chapter 10 Chemical Quantities Worksheet Answers

**Robert C. Smoot, Smoot, Richard G.  
Smith, Jack Price**



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