

1.1 Chemical Reactions

1.1.1 Describing Chemical Reactions

Read to your text to provide balancing chemical equations. Remember that you will be using mass balancing the substances to get the best number of atoms.

1. Balance the equation for the combustion of methane and hydrogen sulfide respectively:



2. Balance the equation for the combustion of ethane and propane respectively:



3. Balance the equation for the synthesis of ammonia and iron sulfide, respectively:



1.1.1.1 Balancing Chemical Equations

4. In chemical reactions matter is not created or destroyed. Change this into an equation: $\underline{Matter is conserved}$

5. The process of combustion is $\underline{exothermic}$

6. In the following reactions, identify which substances are being oxidized or reduced. They are undergoing an $\underline{oxidation}$

7. Write a half-equation for the following reactions:

a. $\underline{Fe} \rightarrow \underline{Fe}^{2+} + \underline{2e^-}$
 b. $\underline{Fe}^{3+} + \underline{e^-} \rightarrow \underline{Fe}^{2+}$

8. Write a balanced equation:



9. In chemical reactions that occur over the electrodes that are the source of electrons the current is produced in \underline{anode}

10. Identify the reactants and products of hydrolysis reactions: a. $\underline{H_2O}$ b. $\underline{OH^-}$



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