

# GAS LAWS AND MOLES

Date: \_\_\_\_\_

Hour: \_\_\_\_\_

## Information: The Gas Constant

We previously examined the fact that  $\frac{PV}{T}$  is a constant as long as the number of moles of a gas remains unchanged.

This means that any time you measure a gas, the pressure times the volume divided by the temperature (in Kelvin) will always equal the same number unless the amount of gas in the container changes. We are now going to explore how the pressure, volume, and temperature depend on the number of moles of gas that are present.

Table 1: Experimental Data for Three Different Samples of Gas That Contain the Same Amount of Gas (8 mol moles of gas is constant)

Gas	# of moles of Gas in container (mol)	Volume of Container (L)	Temperature of Gas (K)	Pressure of Gas (atm)
A	8.0000	10.0	273.0	1.200
B	8.0000	15.0	1200	200.0
C	8.0000	10.0	540.0	6.00

Table 2: Experimental Data for Three Different Samples of Gases That Contain Different Amounts of Gas (8 mol moles of gas is not constant)

Gas	# of moles of Gas in container	Volume of Container (L)	Temperature of Gas (K)	Pressure of Gas (atm)
D	3.5483	40.0	273.0	1.000
E	0.7436	12.5	315	0.5
F	0.00000	10.0	200	1.25

## Critical Thinking Questions

1. Prove that  $\frac{PV}{T}$  is a constant for the data in Table 1, but not a constant for the data in Table 2.

2. Why is  $\frac{PV}{T}$  a constant in Table 1, but not in Table 2?

# Chapter 14 The Behavior Of Gases Worksheet Answers

**Christopher Black, Mark Anestis**



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