

c = specific heat ($J/g^{\circ}C$)
 T = temperature ($^{\circ}C$)

q = heat energy (J)
 m = mass (g)

ΔT = change in

ig 1000 100 007

Name _____ Period _____

Thermal Energy and Heat Unit Test



Part I. For each of the following scenarios: (1 pt. each line)

- Identify the physical change in the scenario (e.g. boiling, vaporization, condensation, melting, or freezing).
- Map out the change in states (example: gas \rightarrow liquid).
- Write whether the process absorbed or released thermal energy.

1. A popsicle melts.

a) _____ b) _____

c) _____

2. Droplets of water form on a cold can of Coca-Cola set out on a hot day.

a) _____ b) _____

c) _____

3. Ice of hands and feet in a refrigerator.

a) _____ b) _____

c) _____

Part II. Specific Heat Calculations Refer to the equation and the symbol legend above.

Hint – use the triangle method. Show your work! (1 pt. each problem.)

1. A 49.3 g sample of solid iron was heated from $16.3^{\circ}C$ to $56.9^{\circ}C$. What was the change in heat energy? (Solve for q (heat energy). Iron has a specific heat value of $449J/g^{\circ}C$.)

$c =$ _____

$m =$ _____

$\Delta T =$ _____

2. A metallurgist heated pellets of aluminum from $21.0^{\circ}C$ to $82.0^{\circ}C$. Aluminum has a specific heat capacity of $902J/g^{\circ}C$. If the energy change was 433 J, what is the mass (g) of the pellets?

$c =$ _____ $\Delta T =$ _____ $q =$ _____

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Judd E. Hollander



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