

Chapter 16 Solubility and Complex Ion Equilibria

Section 16.1: Solubility Equilibria and the Solubility Product

◆ 1. What is Solubility Equilibrium?

- When an **ionic solid** (like AgCl) is put in water, some dissolves into ions.
- At the same time, ions in the water can recombine and form the solid again.
- Eventually, it reaches **equilibrium**: the rate of dissolving = rate of crystallizing.

Example: $AgCl_s \rightleftharpoons Ag^+_{aq} + Cl^-_{aq}$

◆ 2. Solubility Product (K_{sp})

- **K_{sp}** is the equilibrium constant for this dissolving.
- For AgCl: $K_{sp} = [Ag^+][Cl^-]$

👉 Notice: The solid (AgCl) is left out because its concentration is constant.

◆ 3. Solubility vs. K_{sp}

- **Solubility** = how many moles per liter dissolve.
- **K_{sp}** = a number from the ion concentrations at equilibrium.

⚠ Important: A larger K_{sp} doesn't always mean higher solubility because the ion ratios matter.

Example:

- AgCl: $K_{sp} = 1.6 \times 10^{-10}$, solubility $\approx 1.3 \times 10^{-5}$.
 - Ag₃PO₄: $K_{sp} = 1.8 \times 10^{-18}$, solubility $\approx 1.6 \times 10^{-5}$.
Even though Ag₃PO₄ has a smaller K_{sp}, its solubility is close because of the 3:1 ratio of ions.
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Chapter 16 Solubility And Complex Ion Equilibria

Robert K. Wismer



Chapter 16 Solubility And Complex Ion Equilibria:

An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

Problem Exercises for General Chemistry George Gilbert Long, Forrest C. Hentz, 1982 EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS CHANG, 2013-01-07 EBOOK GENERAL CHEMISTRY THE ESSENTIAL CONCEPTS *Principles of Modern Chemistry* David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **General Chemistry - Principles and Structure** James E. Brady, Larry Peck, 1986 **Chemistry** John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant General Chemistry James E. Brady, Gerard E. Humiston, 1982 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics

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web instructions 1 place your biome choose 1 from the above list as a center concept on your map 2 make four branches from the center point for each of the following categories animals plants climate and interesting facts 3 research your biome to include terms and ideas for each of those main categories and link them on your map

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climate biomes may span more than one continent

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web your task is to create an image concept map on a particular biome the concept map should include terms and concepts related to the biome each group will have a different biome and must present the information to the class image can be created using concept mapping software or hand drawn images must be included either digital or cut from

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