

Chapter 11 Nuclear Chemistry

Review

Introduction

Chemical Energy is energy stored in the bonds of atoms and molecules. **Nuclear Energy** is stored through the nucleus of atoms changing their their different nuclei. Nuclear energy is produced by **radioactivity**.

The Nucleus

Composition



Mass Number

- Nucleus is composed of two main particles, protons and neutrons.
- The number of protons is the atomic number.
- The number of protons and neutrons together is the mass number.

Remember that each element has a different mass. Radioactivity is produced because of these atoms.

Radioactivity is the emission of an element's protons, neutrons, electrons, which can be called **alpha**, **beta**, **gamma**, **positron**, **neutron**.

Table 11.1: Radioactive particles, their mass and electric charges			
Particle	Symbol	Mass	Charge
Alpha	${}^4_2\text{He}$	4 amu	+2
Beta	${}^0_{-1}\text{e}$	$\frac{1}{1836}$ amu	-1
Gamma	γ	0	0
Neutron	${}^1_0\text{n}$	1 amu	0
Positron	${}^0_{+1}\text{e}$	$\frac{1}{1836}$ amu	+1

Types of Radioactive Decay

- Alpha decay** is the loss of an α -particle.



- Beta decay** is the loss of a β -particle (a high-speed electron) produced by the nucleus. ${}^A_ZX \rightarrow {}^A_{Z+1}Y + {}^0_{-1}\text{e}$



- Gamma radiation** is the loss of a γ -ray, which is a high-energy γ radiation that is emitted when a radioactive atom loses an α or β particle. ${}^A_ZX \rightarrow {}^A_ZX + \gamma$

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