



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - C_2H_6
 - H_2S
 - CH_4
 - H_2
 - O_2

- Classify each of the following as either ionic or molecular compound. Write the formula for each compound and name the compound.
 - mercury(II) chloride
 - carbon tetrachloride
 - nitrogen dioxide
 - phosphorus pentachloride
 - hydrogen cyanide
 - hydrogen sulfide
 - hydrogen peroxide
 - hydrogen chloride
 - hydrogen bromide
 - hydrogen iodide
 - hydrogen fluoride
 - hydrogen selenide
 - hydrogen telluride
 - hydrogen selenide
 - hydrogen telluride

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride, HF .
- Show covalent bonding in methane, CH_4 .
- Show covalent bonding in carbon dioxide, CO_2 .



- Show covalent bonding in acetylene, C_2H_2 .
- Draw Lewis structures for each of the following molecular compounds. Name each compound.



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sublevel are represented by each of the following quantum numbers?
 - $n=2, l=1, m_l=0, m_s=+\frac{1}{2}$
 - $n=3, l=2, m_l=2, m_s=+\frac{1}{2}$
 - $n=4, l=3, m_l=3, m_s=+\frac{1}{2}$
 - $n=5, l=4, m_l=4, m_s=+\frac{1}{2}$

Chapter 8 Covalent Bonding Test B Answers

Jeffrey C. Pommerville



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