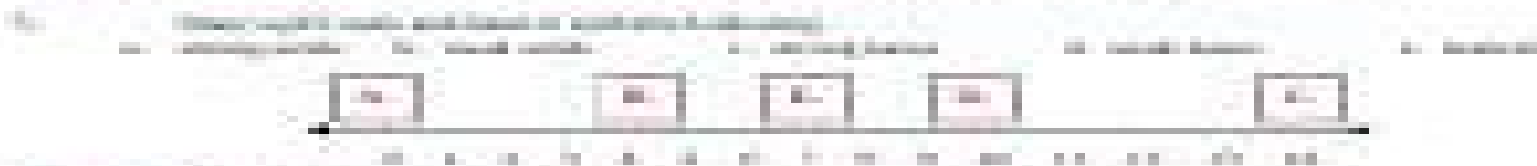


1. Zeichnen Sie die folgenden Komplexe, geben Sie die Oxidationszahl des Zentralatoms an!

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- b) $[PtCl_6]^{2-}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor)
- c) $[PtCl_2(NH_3)_2]$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)
- d) $[PtCl_2(NH_3)_2]^{2+}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)
- e) $[PtCl_2(NH_3)_2]^{2-}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)
- f) $[PtCl_2(NH_3)_2]^{2+}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)
- g) $[PtCl_2(NH_3)_2]^{2-}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)
- h) $[PtCl_2(NH_3)_2]^{2+}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)

- a) $[PtCl_2(NH_3)_2]^{2+}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)
- b) $[PtCl_2(NH_3)_2]^{2-}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)
- c) $[PtCl_2(NH_3)_2]^{2+}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)
- d) $[PtCl_2(NH_3)_2]^{2-}$ → Oxidationszahl Zentralatom: 0 (Platin) & -2 (Chlor) & 0 (Ammoniak)

Zeichnen Sie die folgenden Komplexe, geben Sie die Oxidationszahl des Zentralatoms an!



Zeichnen Sie die folgenden Komplexe, geben Sie die Oxidationszahl des Zentralatoms an!



Zeichnen Sie die folgenden Komplexe, geben Sie die Oxidationszahl des Zentralatoms an!



Zeichnen Sie die folgenden Komplexe, geben Sie die Oxidationszahl des Zentralatoms an!



- a) $[PtCl_2(NH_3)_2]^{2+}$ b) $[PtCl_2(NH_3)_2]^{2-}$
- c) $[PtCl_2(NH_3)_2]^{2+}$ d) $[PtCl_2(NH_3)_2]^{2-}$
- e) $[PtCl_2(NH_3)_2]^{2+}$ f) $[PtCl_2(NH_3)_2]^{2-}$
- g) $[PtCl_2(NH_3)_2]^{2+}$ h) $[PtCl_2(NH_3)_2]^{2-}$

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JR Anderson



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