

MODULE C: SOLIDS, LIQUIDS & GASES

Kinetic Molecular Theory & Phases of Matter

- The different phases of matter can be described based on the Kinetic Molecular Theory

a) According to Kinetic theory, an increase in temperature will cause the average kinetic energy of the particles in the solid, liquid, and gas to increase.

↑ temperature ↑ average kinetic energy

b) When a solid is heated, the temperature is raised and the solid changes to a liquid at its melting point. If the heating is continued, the liquid will change into a gas at its boiling point.

c) The heat energy is changed into the equivalent amount of kinetic energy. Thus, the kinetic energy content of the particles increases in the order:

solid < liquid < gas

- Kinetic Molecular Theory

- 1) Matter is made up of particles
- 2) The particles are always moving

- Kinetic energy

→ energy of movement

- Kinetic Molecular Theory for Gases

a) A gas consists of small particles (atoms or molecules) which have negligible volume (V of gas molecule ≈ 0). This implies that:

$$V_{\text{container}} = V_{\text{molecules}} + V_{\text{free space}} \approx V_{\text{free space}}$$

and that there is plenty of space for compression.

- b) Gas particles are separated relatively far from one another.
- c) Gas particles behave independently from one another. Thus, inter-particle forces of attraction are negligible.
- d) The gas particles are in rapid and continuous motion.
- e) In the gaseous state, the gas particles have vibrational, rotational, and translational motion.
 - 1) Vibrational motion - a particle vibrates about a fixed point
 - 2) Rotational motion - a particle spins about a fixed point
 - 3) Translational motion - a particle moves from one place to another
- f) Pressure of a gas comes from molecules colliding with the walls of the container. Such collisions are perfectly elastic.
- g) Average kinetic energy depends only on the absolute temperature in Kelvin (K)

$$\text{Average Kinetic Energy} = \frac{3}{2} RT, \quad R \rightarrow \text{gas law constant } (0.08206 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})$$

$T \rightarrow$ temperature (in Kelvin)

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Glencoe Science McGraw-Hill Staff, 2001-08

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Table of Contents Chapter Review Part B Concept Answers

1. Understanding the eBook Chapter Review Part B Concept Answers
 - The Rise of Digital Reading Chapter Review Part B Concept Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Review Part B Concept Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Review Part B Concept Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Review Part B Concept Answers

- Personalized Recommendations
 - Chapter Review Part B Concept Answers User Reviews and Ratings
 - Chapter Review Part B Concept Answers and Bestseller Lists
5. Accessing Chapter Review Part B Concept Answers Free and Paid eBooks
 - Chapter Review Part B Concept Answers Public Domain eBooks
 - Chapter Review Part B Concept Answers eBook Subscription Services
 - Chapter Review Part B Concept Answers Budget-Friendly Options
 6. Navigating Chapter Review Part B Concept Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Review Part B Concept Answers Compatibility with Devices
 - Chapter Review Part B Concept Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Review Part B Concept Answers
 - Highlighting and Note-Taking Chapter Review Part B Concept Answers
 - Interactive Elements Chapter Review Part B Concept Answers
 8. Staying Engaged with Chapter Review Part B Concept Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Review Part B Concept Answers
 9. Balancing eBooks and Physical Books Chapter Review Part B Concept Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Review Part B Concept Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter Review Part B Concept Answers
 - Setting Reading Goals Chapter Review Part B Concept Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter Review Part B Concept Answers

- Fact-Checking eBook Content of Chapter Review Part B Concept Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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