

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING**2. PLAN**

- *What step is needed to calculate the concentration of the diluted solution?*

Apply the principle that $\text{volume}_1 \times \text{molarity}_1 = \text{volume}_2 \times \text{molarity}_2$.

$$M_1 V_1 = M_2 V_2 \xrightarrow[\text{substitute each of the known quantities for its symbol and calculate}]{\begin{array}{l} \text{rearrange the equation} \\ M_1 V_1 = M_2 V_2 \\ \text{algebraically to solve for } M_2 \end{array}} M_2 = \frac{M_1 V_1}{V_2}$$

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

3. COMPUTE

Note: Even though molarity is moles per liter, you can use volumes in milliliters along with molarity whenever the units cancel.

$$M_2 = \frac{4.74 \text{ M} \times 50.00 \text{ mL}}{250.00 \text{ mL}} = 0.948 \text{ M}$$

4. EVALUATE

- *Are the units correct?*
- *Is the number of significant figures correct?*
- *Is the answer reasonable?*

Yes; molarity (mol/L) was required.
 Yes; the number of significant figures is correct because the data had a minimum of three significant figures.
 Yes; the computation is the same as $4.74/5$, which is a little less than 1.

PRACTICE

1. Complete the table below by calculating the missing value in each row.

	Molarity of stock solution	Volume of stock solution	Molarity of dilute solution	Volume of dilute solution	
a.	0.500 M KBr	20.00 mL	7 M KBr	100.00 mL	ans: 0.100 M
b.	1.00 M LiOH	? mL	0.075 M LiOH	500.00 mL	ans: 38 mL
c.	7 M HI	5.00 mL	0.0493 M HI	100.00 mL	ans: 0.986 M
d.	12.0 M HCl	0.250 L	1.8 M HCl	? L	ans: 1.7 L
e.	7.44 M NH ₃	? mL	0.093 M NH ₃	4.00 L	ans: 50. mL

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