

## Chemistry - Chapter 7 Study Guide

### Section 7.1:

- *Valence electrons* – electrons in the outer energy level
  - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
  - Look at the group number
    - Group 1A = 1 valence electron
    - Group 7A = 7 valence electrons
    - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
  - Written using the following steps:
    1. Write the element symbol – representing the nucleus and inner electrons
    2. Put one dot for each valence electrons,
    3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
  - Lose electrons to attain a noble gas configuration
  - Written like  $\text{Na}^{1+}$
  - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
  - Gain electrons to attain a noble gas configuration
  - Written like  $\text{P}^{3-}$
  - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

### Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
  - Also called salts
  - Made from a cation and an anion
    - Which means they're made of a metal and a non-metal
- **\*if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain\***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
  - Written like  $\text{Ca}_3\text{P}_2$

# Chemistry Chapter 7 Study Guide Answers

**Ian Pickup**



## **Chemistry Chapter 7 Study Guide Answers:**

*Organic Chemistry Study Guide with Solutions Manual* Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer *Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition* Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines Study Guide, Answer Book to Accompany Organic Chemistry Alan S. Wingrove, Robert L. Caret, 1981 **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994 **Study Guide and Solutions Manual** Neil E. Schore, 2002-08-02 **Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e** Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Study Guide to Accompany Calculus for the Management, Life, and Social Sciences** Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences *Study Guide, Chemical Principles, Fifth Edition, Zumdahl* Paul B. Kelter, Steven S. Zumdahl, 2004-04 Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter *Study Guide for Atoms, Molecules, and Life* Michael S. Matta, 1981 **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 *For All Practical Purposes* Consortium for Mathematics and Its Applications (U.S.), 2003 The sixth edition of the acclaimed classroom favorite offer a number of new features to help instructors strengthen the mathematical literacy of their students **The Technique of Study in First Year Chemistry** Edward Wallace Johnson, 1930 *Organic Chemistry, Study Guide* T. W. Graham Solomons, 1992-01-24 The Fifth Edition has undergone the most extensive revisions of any edition so far Changes include new early chapter on acids and bases in organic chemistry ionic and free radical mechanisms stronger emphasis on the biological environmental medical and industrial applications of organic chemistry as well as on organic synthesis early introduction of the structure and reactivity of carbonyl compounds oxidation reduction reactions and synthesis using Grignard and organolithium reagents

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Solomons,1996      **Organic Chemistry** Francis A. Carey,2000 A Market Leading Traditional Approach to Organic Chemistry Throughout all seven editions Organic Chemistry has been designed to meet the needs of the mainstream two semester undergraduate organic chemistry course This best selling text gives students a solid understanding of organic chemistry by stressing how fundamental reaction mechanisms function and reactions occur With the addition of handwritten solutions new cutting edge molecular illustrations updated spectroscopy coverage seamless integration of molecular modeling exercises and state of the art multimedia tools the 7th edition of Organic Chemistry clearly offers the most up to date approach to the study of organic chemistry      **Study Guide** Joan I. Senyk,Darrell D. Ebbing,1984

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