

Answer all the questions
in

1.

A is a solution containing 4.00 g of the acid, HX, per dm^3 .
B contains 1.00 g sodium hydroxide per 250 cm^3 of solution.

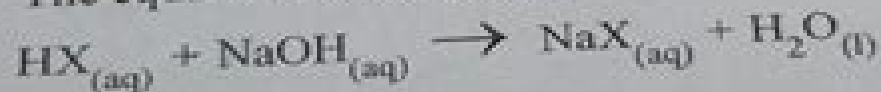
A was titrated with 20.0 cm^3 portion of **B** using methyl orange as indicator and the following results were obtained.

Burette Readings / cm^3	1	2	3
Final reading	23.10	24.60	46.30
Initial reading		1.50	
Volume of A used	23.10		23.20

- (a)
- Copy and complete the table.
 - Calculate the average volume of **A** used.
 - State the colour of the indicator:
 - in solution **B** before the titration;
 - at the end-point.

- (b) From the information provided, calculate the:
- concentration of **B** in mol dm^{-3} ;
 - concentration of **A** in mol dm^{-3} ;
 - molar mass of the acid HX.

The equation for the reaction is:



$$[\text{H} = 1.0, \text{O} = 16.0, \text{Na} = 23.0]$$

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