

# Chemistry

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(Chapter 4) (The d and f block Elements) (Exercise)  
(Class 12)

## Question 4.23:

Which metal in the first series of transition metals exhibits +1 oxidation state most frequently and why?

### Answer 4.23:

In the first transition series, Cu exhibits +1 oxidation state very frequently. It is because Cu (+1) has an electronic configuration of  $[Ar]3d^{10}$ . The completely filled  $d$ -orbital makes it highly stable.

## Question 4.24:

Calculate the number of unpaired electrons in the following gaseous ions:  $Mn^{3+}$ ,  $Cr^{3+}$ ,  $V^{3+}$  and  $Ti^{3+}$ . Which one of these is the more stable in aqueous solution?

### Answer 4.24:

	Gaseous ions	Number of unpaired electrons
(i)	$Mn^{3+}$ , $[Ar]3d^4$	4
(ii)	$Cr^{3+}$ , $[Ar]3d^3$	3
(iii)	$V^{3+}$ , $[Ar]3d^2$	2
(iv)	$Ti^{3+}$ , $[Ar]3d^1$	1

$Cr^{3+}$  is the most stable in aqueous solutions owing to a  $t_{2g}^3$  configuration.

## Question 4.25:

Give examples and suggest reasons for the following features of the transition metal chemistry:

(i) The lowest oxide of transition metal is basic, the highest is amphoteric/acidic.

(ii) A transition metal exhibits highest oxidation state in oxides and fluorides.

(iii) The highest oxidation state is exhibited in oxoanions of a metal.

### Answer 4.25:

(i) In the case of a lower oxide of a transition metal, the metal atom has a low oxidation state. This means that some of the valence electrons of the metal atom are not involved in bonding. As a result, it can donate electrons and behave as a base.

On the other hand, in the case of a higher oxide of a transition metal, the metal atom has a high oxidation state. This means that the valence electrons are involved in bonding and so, they are unavailable. There is also a high effective nuclear charge. As a result, it can accept electrons and behave as an acid.

For example,  $Mn^{II}O$  is basic and  $Mn^{VII}O_7$  is acidic.

(ii) Oxygen and fluorine act as strong oxidising agents because of their high electro negativities and small sizes. Hence, they bring out the highest oxidation states from the transition metals. In other words, a transition metal exhibits higher oxidation states in oxides and fluorides. For example, in  $OsF_6$  and  $V_2O_5$ , the oxidation states of Os and V are +6 and +5 respectively.

(iii) Oxygen is a strong oxidising agent due to its high electronegativity and small size. So, oxo-anions of a metal have the highest oxidation state. For example, in  $MnO_4^-$ , the oxidation state of Mn is +7.

## Question 4.26:

Indicate the steps in the preparation of:

(i)  $K_2Cr_2O_7$  from chromite ore.

(ii)  $KMnO_4$  from pyrolusite ore.

### Answer 4.26:

(i) Potassium dichromate ( $K_2Cr_2O_7$ ) is prepared from chromite ore ( $FeCr_2O_4$ ) in the following steps:

Step (i): Preparation of sodium chromate



Step (ii): Conversion of sodium chromate into sodium dichromate



Step (iii): Conversion of sodium dichromate to potassium dichromate



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## **Table of Contents Chemistry Chapter 4 Answers**

1. Understanding the eBook Chemistry Chapter 4 Answers
  - The Rise of Digital Reading Chemistry Chapter 4 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 4 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 4 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 4 Answers
  - Personalized Recommendations
  - Chemistry Chapter 4 Answers User Reviews and Ratings
  - Chemistry Chapter 4 Answers and Bestseller Lists
5. Accessing Chemistry Chapter 4 Answers Free and Paid eBooks
  - Chemistry Chapter 4 Answers Public Domain eBooks
  - Chemistry Chapter 4 Answers eBook Subscription Services
  - Chemistry Chapter 4 Answers Budget-Friendly Options

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  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 4 Answers Compatibility with Devices
  - Chemistry Chapter 4 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 4 Answers
  - Highlighting and Note-Taking Chemistry Chapter 4 Answers
  - Interactive Elements Chemistry Chapter 4 Answers
8. Staying Engaged with Chemistry Chapter 4 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 4 Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 4 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 4 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 4 Answers
  - Setting Reading Goals Chemistry Chapter 4 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 4 Answers
  - Fact-Checking eBook Content of Chemistry Chapter 4 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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