

**Chemical Bonding****SECTION 3**

**SHORT ANSWER** Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
  - formula unit.
  - molecule.
  - crystal.
  - atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
  - molecules.
  - positive ions.
  - dipoles.
  - negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
  - higher in potential energy.
  - lower in potential energy.
  - equal in potential energy.
  - unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
  - Compound A is not an ionic compound.
  - It will be more difficult to break the bonds in compound A than those in compound B.
  - Compound B has larger crystals than compound A.
  - Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
  - stronger than the attractive forces among formula units in ionic bonding.
  - weaker than the attractive forces among formula units in ionic bonding.
  - approximately equal to the attractive forces among formula units in ionic bonding.
  - equal to zero.
- Describe the force that holds two ions together in an ionic bond.

**Very strong overall attraction between positive and negative charges**

- What type of energy best represents the strength of an ionic bond?

**Lattice energy**

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

# Chemistry Chapter 6 Review Answers

**Albert A Gayle**



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## **Table of Contents Chemistry Chapter 6 Review Answers**

1. Understanding the eBook Chemistry Chapter 6 Review Answers
  - The Rise of Digital Reading Chemistry Chapter 6 Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 6 Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 6 Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 6 Review Answers
  - Personalized Recommendations
  - Chemistry Chapter 6 Review Answers User Reviews and Ratings
  - Chemistry Chapter 6 Review Answers and Bestseller Lists
5. Accessing Chemistry Chapter 6 Review Answers Free and Paid eBooks
  - Chemistry Chapter 6 Review Answers Public Domain eBooks
  - Chemistry Chapter 6 Review Answers eBook Subscription Services
  - Chemistry Chapter 6 Review Answers Budget-Friendly Options

6. Navigating Chemistry Chapter 6 Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 6 Review Answers Compatibility with Devices
  - Chemistry Chapter 6 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 6 Review Answers
  - Highlighting and Note-Taking Chemistry Chapter 6 Review Answers
  - Interactive Elements Chemistry Chapter 6 Review Answers
8. Staying Engaged with Chemistry Chapter 6 Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 6 Review Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 6 Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 6 Review Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 6 Review Answers
  - Setting Reading Goals Chemistry Chapter 6 Review Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 6 Review Answers
  - Fact-Checking eBook Content of Chemistry Chapter 6 Review Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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