

Chapter 6

Equilibrium

Question 6.47:

It has been found that the pH of a 0.01M solution of an organic acid is 4.15. Calculate the concentration of the anion, the ionization constant of the acid and its pK_a .

Answer 6.47:

Let the organic acid be HA.



Concentration of HA = 0.01 M pH = 4.15

$$-\log[H^+] = 4.15$$

$$[H^+] = 7.08 \times 10^{-5}$$

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

Now,

$$[H^+] = [A^-] = 7.08 \times 10^{-5}$$

$$[HA] = 0.01$$

Then,

$$K_a = \frac{(7.08 \times 10^{-5})(7.08 \times 10^{-5})}{0.01} = 5.01 \times 10^{-7}$$

$$pK_a = -\log K_a = -\log(5.01 \times 10^{-7}) = 6.3001$$

Question 6.48:

Assuming complete dissociation, calculate the pH of the following solutions:

(i) 0.003 M HCl

(ii) 0.005 M NaOH

(iii) 0.002 M HBr

(iv) 0.002 M KOH

Answer 6.48:

(i) 0.003M HCl:



Since HCl is completely ionized,

$$[H_3O^+] = [HCl] = 0.003$$

Now,

$$pH = \log[H_3O^+] = -\log(.003) = 2.52$$

Hence, the pH of the solution is 2.52.

Chemistry Chapter 6 Solutions Manual

Robert F. Francis



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General, Organic, and Biological Chemistry Dorothy M. Feigl, John William Hill, 1986

Chemistry George M. Bodner, Harry L. Pardue, 1989-01-23 A modern experimental approach to first year chemistry This unique introductory account employs experimental observations to construct the principles of general chemistry An early introduction to observable descriptive chemistry lays the basis for the well developed exposition that follows

Organic Chemistry Stanislaw Skonieczny, George H. Schmid, 1996

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