

Periodic Table Trend Activities:



Purpose: The Periodic Table is a useful tool that helps us predict properties of various elements. In this activity, we will look at 3 properties:

1. **Atomic radius:** The distance from the atomic nucleus to the outmost occupied orbital in a atom
2. **Electronegativity:** The ability of an atom to attract a bonded pair of electrons from another atom
3. **Ionization Energy:** The energy needed to remove an electron from a neutral gaseous atom
4. **Electron Affinity:** The energy released when an electron is added to a neutral gaseous atom

Procedure: Click on the [Cool Periodic Table link](#) (click on the element symbol). Record the atomic radius, electronegativity, and first ionization energy for each element in the table below:

<http://www.lyncsburg.net/hhs/chemistry/trends/>

Data:

| Element | Atomic Radius | Electronegativity | 1 st Ionization Energy |
|---------|---------------|-------------------|-----------------------------------|
| Li | | | |
| Be | | | |
| B | | | |
| Na | | | |
| Mg | | | |
| Al | | | |
| C | | | |
| Si | | | |
| N | | | |
| P | | | |

Discussion:

1. What appears to be the trend in atomic radius as you move from left to right in a row?
2. What appears to be the trend in atomic radius as you move down a column?

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M Tight



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