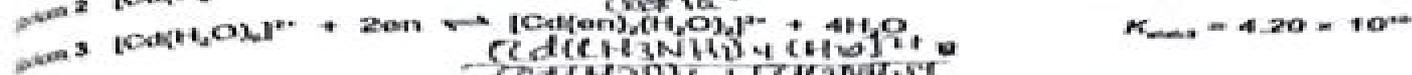
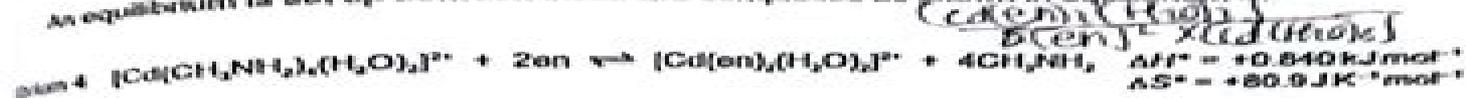


(b) Cadmium ions form complexes with methylamine, CH_3NH_2 , and with 1,2-diaminoethane, $\text{H}_2\text{NCH}_2\text{CH}_2\text{NH}_2$, as shown in equilibria 2 and 3. 1,2-diaminoethane is shown as en.



An equilibrium is set up between these two complexes as shown in equilibrium 4.



(i) K_{eq} is the equilibrium constant for equilibrium 4.

Write an expression for K_{eq} in terms of $K_{\text{stab}2}$ and $K_{\text{stab}3}$.

$$K_{\text{eq}} = \frac{[\text{CH}_3\text{NH}_2]^4 \times [\text{Cd}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}}{[\text{en}]^2 \times [\text{Cd}(\text{CH}_3\text{NH}_2)_4(\text{H}_2\text{O})_2]^{2+}} = \frac{K_{\text{stab}3}}{K_{\text{stab}2}} \quad [1]$$

(ii) Calculate the value of the standard Gibbs free energy change, ΔG° , for equilibrium 4 at 298 K.

$$\Delta G = \Delta H - T\Delta S = +0.840 - 298 \times (80.9 \times 10^{-3})$$

$$\Delta G^\circ = -23.3 \text{ kJ mol}^{-1} \quad [2]$$

(iii) State how the value of ΔG° changes as the temperature increases. Explain your answer.

As temp increases, ΔG value gets more negative as value for $(T\Delta S)$ in the eqn increases. [1]

[Total: 9]

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Table of Contents Chemistry Paper 42 November 2002 Mark Scheme

1. Understanding the eBook Chemistry Paper 42 November 2002 Mark Scheme
 - The Rise of Digital Reading Chemistry Paper 42 November 2002 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Paper 42 November 2002 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Paper 42 November 2002 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Paper 42 November 2002 Mark Scheme
 - Personalized Recommendations
 - Chemistry Paper 42 November 2002 Mark Scheme User Reviews and Ratings

- Chemistry Paper 42 November 2002 Mark Scheme and Bestseller Lists
- 5. Accessing Chemistry Paper 42 November 2002 Mark Scheme Free and Paid eBooks
 - Chemistry Paper 42 November 2002 Mark Scheme Public Domain eBooks
 - Chemistry Paper 42 November 2002 Mark Scheme eBook Subscription Services
 - Chemistry Paper 42 November 2002 Mark Scheme Budget-Friendly Options
- 6. Navigating Chemistry Paper 42 November 2002 Mark Scheme eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Paper 42 November 2002 Mark Scheme Compatibility with Devices
 - Chemistry Paper 42 November 2002 Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Paper 42 November 2002 Mark Scheme
 - Highlighting and Note-Taking Chemistry Paper 42 November 2002 Mark Scheme
 - Interactive Elements Chemistry Paper 42 November 2002 Mark Scheme
- 8. Staying Engaged with Chemistry Paper 42 November 2002 Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Paper 42 November 2002 Mark Scheme
- 9. Balancing eBooks and Physical Books Chemistry Paper 42 November 2002 Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Paper 42 November 2002 Mark Scheme
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Paper 42 November 2002 Mark Scheme
 - Setting Reading Goals Chemistry Paper 42 November 2002 Mark Scheme
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Paper 42 November 2002 Mark Scheme
 - Fact-Checking eBook Content of Chemistry Paper 42 November 2002 Mark Scheme
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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