

## Ionic

Name: \_\_\_\_\_

Date: \_\_\_\_\_

Hour: \_\_\_\_\_

## Bonding

**Information: Naming Ions**

To write an ion, you write the symbol of the atom and put the charge in the upper right corner. Consider the following examples:  $\text{Al}^{3+}$ ,  $\text{O}^{2-}$ ,  $\text{Mg}^{2+}$ . You should verify that each of the charges is correct.

Positive and negative ions are named differently. Positive ions retain the same name as the parent atom. For example,  $\text{Al}^{3+}$  is called the "aluminum ion" and  $\text{Mg}^{2+}$  is called the "magnesium ion." Negative ions are named a little differently. For negative ions, you change the ending of the name to "-ide". Therefore,  $\text{O}^{2-}$  is named oxide and  $\text{P}^{3-}$  is named phosphide.

**Critical Thinking Questions**

- Sulfur forms a negatively charge ion. Is the name of the ion sulfur or sulfide? Explain why.
- Sodium forms a positive ion. Is the name of the ion sodium or sodide? Explain why.
- Write the symbol (including the charge) AND the name for each of the ions for each of the following. The first is done for you.

a) Cl

b) Ca

c) N

d) K

e) S

f) B

g) P

$\text{Cl}^{-1}$   
chloride

**Information: Ionic Bonding and Formulas**

There are two ways in which atoms can "bond" to each other and form a compound. The means of bonding that we will consider now is called ionic bonding.

Ionic bonding occurs between a metal and a nonmetal. As you know, opposite charges attract. Ionic bonding is when two ions of opposite charge attract and bond to each other forming an ionic compound.

Consider the following examples of formulas for ionic compounds:

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**Michal Rosen-Zvi**



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