

Experiment 11A: The Periodic Table

Procedure

Part 1 Electronic Configuration and the Periodic Table

1. Obtain a sheet of paper.

Part 2 Atomic Number, Atomic Symbol, Electronic Configuration and the Periodic Table

1. Obtain a sheet of paper (graph paper preferred).

Part 3 Electronic Configuration and Energy Levels (Periodic Table)

1. Obtain a sheet of paper.

Background

1. Examine the periodic table of elements and compare it to Figure 11.1. What relationship can be seen between an element's placement in the periodic table and its electronic configuration?

Elements with similar electronic configurations are located in similar places on the periodic table. Group 1 elements are alkali metals, group 2 elements are alkaline earth metals, and group 18 elements are noble gases.

2. Examine the graph of atomic radii plotted against increasing atomic number (Figure 11.4). What is particularly noteworthy about it? Is it similar to the periodic table?

Atomic radii are plotted against atomic number on the periodic table graph. Examining the graph shows that the radii generally increase as you go down the periodic table. The atomic radii generally decrease as you go across the periodic table. However, there is a sharp drop in atomic radii between group 2 elements and group 13 elements on the periodic table.

3. Which group of elements has the largest atomic radii? Which group has the smallest?

Examine the graph of ionization energy plotted against increasing atomic number (Figure 11.5). What is particularly noteworthy about it? Is it similar to the periodic table?

4. Examine the graph of the first ionization energy plotted against increasing atomic number (Figure 11.6). What is particularly noteworthy about it? Is it similar to the periodic table?

Ionization energy is plotted against atomic number on the periodic table graph. The graph shows that ionization energy generally increases as you go across the periodic table. However, there is a sharp drop in ionization energy between group 2 elements and group 13 elements on the periodic table.

Laboratory Experiment 11a Answers Periodic Table

JR Anderson



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