

Format for B.Tech Project Synopsis

Title page:

1. Name of Student
2. PTU Roll No.
3. Class Roll No.
4. Present official Address with E-mail, telephone No.
5. Branch
6. Batch
7. Proposed Topic

Introduction (should not exceed 3 pages)

The introduction part will include the brief introduction about the project to be developed, technology used, field of project (if specialized one), any special technical terms about the project.

Literature Survey: (should not exceed 2 pages)

This will describe the previous work which is done by the various authors/developers related to the project and already running project at any location.

Methodology/ Planning of work (should not exceed 1 page)

Methodology will include the steps to be followed to complete the project during the project development.

Facilities required for proposed work

Software/Hardware required for the development of the project.

References

Here specify the description of the study material referred for the development of the project.

Project Synopsis Format Mechanical Engineering Notes

Chao Zhang



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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} \rightarrow 2\text{HCN} + 2\text{OH}^{-}$ $\text{H}_2\text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-} \rightarrow 2\text{H}_2\text{O}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-} \rightarrow 2\text{H}_2\text{O} + 2\text{OH}^{-}$ 3 the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a

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